



NATIONAL  
EXAM CERTIFICATE

GRADE 11

NOVEMBER 2020

PHYSICAL SCIENCES P1  
(CHEMISTRY),  
EXEMPLAR

MARKS  
100



The question paper consists of 10 pages, including this cover page.  
There are 10 questions in this question paper.

# QUESTION 3

## PHYSICAL SCIENCES PAPER 2 INTERMOLECULAR FORCES

**GRADE 11**

**NOVEMBER 2020**

QUESTION 3 MARKS AVAILABLE 10

QUESTION 3 TIME ALLOWED 30 minutes

- 3.1 The following table shows the interaction energy observed during an intermolecular force experiment for halogen molecules in group 17. Complete the table and calculate the mean interaction energy of the four molecules.

Molecule	Interaction energy (kJ/mol)	Percentage (%)
Cl <sub>2</sub>	-100	50
I <sub>2</sub>	-100	50

- 3.2 Calculate the interaction energy between two molecules of iodine (I<sub>2</sub>) and the polar nature of their molecule.
- 3.3 State what is the maximum size of the van der Waals radius of a molecule.
- 3.4 Define the term 'polar molecule'.
- 3.5 Explain how the surface tension of liquids is affected by the presence of solutes.
- 3.6 Calculate the surface tension of water from a surface tension test that uses a capillary tube if the radius of curvature of the meniscus is 0.01 m, and state the units from a similar test where the radius of curvature of the meniscus is 0.001 m.
- 3.7 Explain why the boiling point of chlorine is higher than that of bromine at 20 °C. State the effect of the Fajan rule on the interaction forces between molecules.
- 3.8 Calculate the total interaction of two molecules of iodine (I<sub>2</sub>) if the distance between them is 0.1 nm.



- 3.9 Explain which molecule has the higher density?
- 3.10 Explain the effect of temperature on the viscosity of a non-Newtonian fluid.

QUESTION 3  
ANSWER



Virtual Tutorship

# Physical Science Paper 2 Intermolecular Forces

## Investigation

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