

# Physics

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(Chapter 11) (Thermal Properties of Matter) (Exercises)  
(Class 11)

$$\text{Useful Energy, } \Delta Q = \frac{50}{100} \times 1.5 \times 10^6 = 7.5 \times 10^5 \text{ J}$$

$$\text{But } \Delta Q = mc\Delta T$$

$$\therefore \Delta T = \frac{\Delta Q}{mc} = \frac{7.5 \times 10^5}{8 \times 10^3 \times 0.91} = 103^\circ\text{C}$$

Therefore, in 2.5 minutes of drilling, the rise in the temperature of the block is  $103^\circ\text{C}$ .

## Question 11.13:

A copper block of mass 2.5 kg is heated in a furnace to a temperature of  $500^\circ\text{C}$  and then placed on a large ice block. What is the maximum amount of ice that can melt? (Specific heat of copper =  $0.39 \text{ J g}^{-1} \text{ K}^{-1}$ ; heat of fusion of water =  $335 \text{ J g}^{-1}$ )

## Answer 11.13:

Mass of the copper block,  $m = 2.5 \text{ kg} = 2500 \text{ g}$

Rise in the temperature of the copper block,  $\Delta\theta = 500^\circ\text{C}$

Specific heat of copper,  $C = 0.39 \text{ J g}^{-1} \text{ }^\circ\text{C}^{-1}$

Heat of fusion of water,  $L = 335 \text{ J g}^{-1}$

The maximum heat the copper block can lose,  $Q = mC\Delta\theta = 2500 \times 0.39 \times 500 = 487500 \text{ J}$

Let  $m_1 \text{ g}$  be the amount of ice that melts when the copper block is placed on the ice block.

The heat gained by the melted ice,  $Q = m_1 L$

$$\therefore m_1 = \frac{Q}{L} = \frac{487500}{335} = 1455.22 \text{ g}$$

Hence, the maximum amount of ice that can melt is 1.45 kg.

## Question 11.14:

In an experiment on the specific heat of a metal, a 0.20 kg block of the metal at  $150^\circ\text{C}$  is dropped in a copper calorimeter (of water equivalent 0.025 kg) containing  $150 \text{ cm}^3$  of water at  $27^\circ\text{C}$ . The final temperature is  $40^\circ\text{C}$ . Compute the specific heat of the metal. If heat losses to the surroundings are not negligible, is your answer greater or smaller than the actual value for specific heat of the metal?

## Answer 11.14:

Mass of the metal,  $m = 0.20 \text{ kg} = 200 \text{ g}$

Initial temperature of the metal,  $T_1 = 150^\circ\text{C}$  and Final temperature of the metal,  $T_2 = 40^\circ\text{C}$

Calorimeter has water equivalent of mass,  $m' = 0.025 \text{ kg} = 25 \text{ g}$

Volume of water,  $V = 150 \text{ cm}^3$

Mass ( $M$ ) of water at temperature  $T = 27^\circ\text{C}$ :  $150 \times 1 = 150 \text{ g}$

Fall in the temperature of the metal:  $\Delta T = T_1 - T_2 = 150 - 40 = 110^\circ\text{C}$

Specific heat of water,  $C_w = 4.186 \text{ J/g/}^\circ\text{K}$  and Specific heat of the metal =  $C$

Heat lost by the metal,  $\theta = mC\Delta T \dots (i)$

Rise in the temperature of the water and calorimeter system:

$$\Delta T' = 40 - 27 = 13^\circ\text{C}$$

Heat gained by the water and calorimeter system:  $\Delta\theta' = m_1 C_w \Delta T' = (M + m') C_w \Delta T' \dots (ii)$

Heat lost by the metal = Heat gained by the water and calorimeter system

$$mC\Delta T = (M + m') C_w \Delta T' \Rightarrow 200 \times C \times 110 = (150 + 25) \times 4.186 \times 13$$

$$\therefore C = \frac{175 \times 4.186 \times 13}{110 \times 200} = 0.43 \text{ J g}^{-1} \text{ K}^{-1}$$

If some heat is lost to the surroundings, then the value of  $C$  will be smaller than the actual value.

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