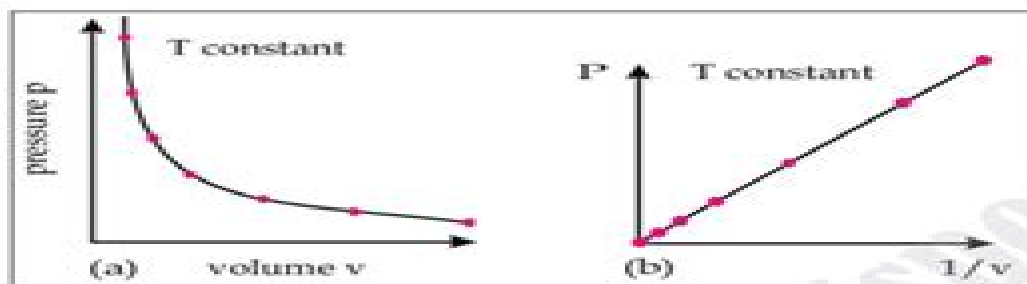


- (1) If the pressure of the gas is double volume becomes half. If pressure increases by three times then the volume becomes one-third and so on.
- (2) The graph between "p" and "V" between "p and $\frac{1}{V}$ " is shown below:



- (3) The graph between "p" and "V" shows that if pressure increases then volume decreases and vice-versa.
- (4) The graph between "p" and " $1/V$ " shows a straight line passing through the origin.
- (5) At constant temperature the product of pressure and volume is constant. i.e.

$$pV = \text{constant}$$
- (6) Using the above result, at a constant temperature, we can write:
 initial pressure \times initial volume = final pressure \times final volume

$$p_1V_1 = p_2V_2$$

Thus Robert Boyle concluded his Law known as "Boyle's Law" which states that:

Statement of Boyle's Law:

The volume of a fixed mass of a gas is inversely proportional to its pressure, provided its temperature remains constant.

Q.23: Write down the application of Boyle's Law.

Ans: Applications of (p-V) relationship of gas / Boyle's Law:

Some applications of pressure-volume (p-V) relationship of a gas i.e. Boyle's Law is given below:

Aerosols, such as spray paints, use Boyle's Law in their working mechanism.

When we breathe, our diaphragm moves downward, increasing the volume of the Lungs. This causes the pressure inside the Lungs to be less than the outside pressure so air rushes in.

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J Dewey



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