

Chapter 1: Answers to Review Problems

Section 1.1

- 1) The energy of electromagnetic radiation is inversely proportional to its wavelength; the shorter the wavelength, the higher the energy.
The frequency of a wave refers to the number of times a wave repeats itself in a given amount of time. The shorter the wavelength, the higher the frequency.
Energy and frequency are directly proportional; the higher the frequency, the higher the energy.
- 2) Each of these series refers to a group of spectral lines, each line corresponding to a specific wavelength of electromagnetic radiation, and therefore, a specific energy of radiation.
The Bohr model states that normally the electron in hydrogen occupies the lowest energy level and this is called the ground state. If the electron absorbs energy, it is promoted to a higher energy level, known as an excited state. As the excited state is not stable, the electron will eventually return to a lower energy level, emitting the absorbed energy as light. The energy difference between the excited state and the ground state corresponds to light of a certain wavelength.
The Balmer series refers to wavelengths of light emitted when an electron falls from a higher energy level to $n = 2$; these wavelengths fall in the visible region of the spectrum.
The Lyman series refers to wavelengths of light emitted when an electron falls from a higher energy level to $n = 1$; these wavelengths fall in the ultraviolet region of the spectrum.
The Paschen series refers to wavelengths of light emitted when an electron falls from a higher energy level to $n = 3$; these wavelengths fall in the infrared region of the spectrum.
- 3) The $\Psi_2(x)$ wave has a shorter wavelength than that of the $\Psi_1(x)$ wave. Therefore, the energy of the $\Psi_2(x)$ wave is greater than that of the $\Psi_1(x)$ wave. Also, the greater the number of nodes, the higher the energy.
- 4) **Pauli's Exclusion Principle** states that no two electrons in an atom will have the same set of four quantum numbers. The valence shell configuration of Carbon is $2s^2 2p^2$
For the two electrons in the $2s$:
Electron # 1: $n = 2, l = 0, m_l = 0, m_s = +\frac{1}{2}$
Electron # 2: $n = 2, l = 0, m_l = 0, m_s = -\frac{1}{2}$
These two electrons have different values for m_s , so the two electrons have a different set of 4 quantum numbers.

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