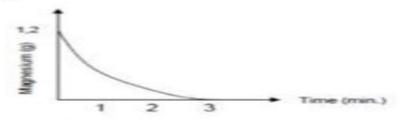
Grade 11 Physical Science M...

records the change in the mass of magnesium as the reaction proceeds at regular intervals. The balanced chemical equation for the reaction is:

$$Mg(s) + 2HCl(aq) \rightarrow MgCl_2(aq) + H_2(g)$$

The change in the mass of magnesium during the reaction is shown on the graph below.



- 6.2.1 Identify the limiting agent in this reaction. Give a reason for the answer.
- 6.2.2 Calculate the number of moles of unreacted hydrochloric acid in the test tube after 3 minutes.

Hide Answers

QUESTION/VRAAG 6

- 6.1.1 2H₂O₂ (aq) ✓ → 2H₂O (ℓ) + O₂ (g) ✓
- 6.1.2 The catalyst lowers the activation energy of the reaction
 Accept: catalyst speeds up the reaction
 In Katalisator vertiage die aktiveringsenergie van die reaksie
 - 'n Katalisator verlaag die aktiveringsenergie van die reaksie 🗸 🗸 Aanvaar: katalisator laat die reaksie vinniger plaasvind
- 6.1.3 OPTION 1/OPSIE 1

$$0,049 = \frac{m}{34} \checkmark$$

OPTION 2/OPSIE 2
From the balanced equation:
Vanaf gebalanseerde
vergelyking:

68g $H_2O_2 \rightarrow 24,45 \text{ dm}^3 O_2 \checkmark \checkmark$ $X \text{ g } H_2O_2 \rightarrow 600 \times 10^{-3} \text{ dm}^3 \checkmark$ $X = 68 \times 0.6 \checkmark$ (2)

(7) [19]

(2)

$$X = \frac{68 \times 0.6}{24.45}$$

 $X = 1.67 g$

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- $1. \ \ Understanding \ the \ eBook \ Physical \ Sciences \ Grade \ 11 \ Exampler \ 2014$
 - The Rise of Digital Reading Physical Sciences Grade 11 Exampler 2014
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Physical Sciences Grade 11 Exampler 2014
 - Exploring Different Genres
 - o Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Physical Sciences Grade 11 Exampler 2014
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Physical Sciences Grade 11 Exampler 2014
 - Personalized Recommendations
 - Physical Sciences Grade 11 Exampler 2014 User Reviews and Ratings

- Physical Sciences Grade 11 Exampler 2014 and Bestseller Lists
- 5. Accessing Physical Sciences Grade 11 Exampler 2014 Free and Paid eBooks
 - Physical Sciences Grade 11 Exampler 2014 Public Domain eBooks
 - Physical Sciences Grade 11 Exampler 2014 eBook Subscription Services
 - Physical Sciences Grade 11 Exampler 2014 Budget-Friendly Options
- 6. Navigating Physical Sciences Grade 11 Exampler 2014 eBook Formats
 - o ePub, PDF, MOBI, and More
 - Physical Sciences Grade 11 Exampler 2014 Compatibility with Devices
 - Physical Sciences Grade 11 Exampler 2014 Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - o Adjustable Fonts and Text Sizes of Physical Sciences Grade 11 Exampler 2014
 - Highlighting and Note-Taking Physical Sciences Grade 11 Exampler 2014
 - Interactive Elements Physical Sciences Grade 11 Exampler 2014
- 8. Staying Engaged with Physical Sciences Grade 11 Exampler 2014
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Physical Sciences Grade 11 Exampler 2014
- 9. Balancing eBooks and Physical Books Physical Sciences Grade 11 Exampler 2014
 - \circ Benefits of a Digital Library
 - Creating a Diverse Reading Collection Physical Sciences Grade 11 Exampler 2014
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Physical Sciences Grade 11 Exampler 2014
 - Setting Reading Goals Physical Sciences Grade 11 Exampler 2014
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Physical Sciences Grade 11 Exampler 2014
 - Fact-Checking eBook Content of Physical Sciences Grade 11 Exampler 2014
 - Distinguishing Credible Sources

- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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