## A. Covalent & Ionic Compounds/Bonds

When an atom loses an electron it becomes a \_\_\_\_positive ion \_\_\_\_\_ (cation),

and when it gains an electron it becomes a \_\_\_\_negative ion \_\_\_\_\_ (anion).

An element in the halogen group has \_7\_\_ valence electrons, and wants to \_\_\_gain 1 more to have S. This is because of the \_\_octet\_\_\_\_ rule.

3. An \_ionic\_\_\_\_ bond is the attraction between oppositely charged ions, AKA a metal and a nonmetal .

In a covalent bond, atoms \_\_\_\_share\_\_ electrons, whereas in an ionic bond, atoms transfer electrons.

5. (Honors only) When electrons are shared unequally in a covalent bond it is \_\_\_\_\_\_\_\_\_, and when they're shared equally it is \_\_\_\_\_\_\_nonpolar\_\_\_\_\_\_. You can predict if a compound is ionic, polar covalent or nonpolar covalent by subtracting the \_\_\_\_\_\_electronegativities\_\_\_\_\_\_.

 An atom can have up to \_8\_\_\_ valence electrons. The family (column) on the periodic table that all have 8 valence electrons is called \_\_\_\_ noble gases\_\_\_\_.

7. Atoms can gain, lose, or share \_\_\_valence electrons\_\_\_\_ to become more stable.

## B. Formulas for Ionic compounds - Fill in the chart below

and the second s	Magnesium	Fluorine	Nitrogen	Lithium
# OF VALENCE ELECTRONS	2	7	5	1
OXIDATION #	+2	-1	-3	+1

2. What is the **chemical formula** for each of the following ionic compounds? (hint: you can either draw the Lewis structures and use arrows to show the transfer of electrons, OR use the periodic table to find their oxidation numbers. They will combine so that their oxidation numbers cancel out to zero. Use the criss cross method if their oxidation numbers do not cancel.)

- a) Potassium and Chlorine \_\_\_\_KCl\_\_\_
- b) Calcium and Bromine \_\_\_\_CaBr2\_\_\_\_
- c) Magnesium and oxygen \_\_\_MgO\_\_\_

# **Physical Science Study Guide Chemical Bonding**

**DJ Losen** 

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