

Worksheet: Periodic Trends

Name _____
Period _____

1. Which statement best describes Group 2 elements as they are considered in order from top to bottom of the Periodic Table?
(A) The number of principal energy levels increases, and the number of valence electrons increases.
(B) The number of principal energy levels increases, and the number of valence electrons remains the same.
(C) The number of principal energy levels remains the same, and the number of valence electrons increases.
(D) The number of principal energy levels remains the same, and the number of valence electrons decreases.
2. What is the total number of valence electrons in an atom of boron in the ground state?
(A) 1 (C) 3
(B) 7 (D) 5
3. What is the total number of valence electrons in an atom of xenon, Xe?
(A) 6 (C) 8
(B) 2 (D) 18
4. The elements calcium and strontium have similar chemical properties because they both have the same
(A) atomic number
(B) mass number
(C) number of valence electrons
(D) number of completely filled sublevels
5. On the Periodic Table of the Elements, all the elements within Group 16 have the same number of
(A) valence electrons (C) protons
(B) energy levels (D) neutrons
6. An element with a partially filled *d* sublevel in the ground state is classified as
(A) a halogen (C) an alkali metal
(B) a transition metal (D) an alkaline earth metal
7. Which electron configuration represents a transition element?
(A) $1s^2 2s^2 2p^2$ (C) $[Ar]3d^5 4s^2$
(B) $[Ne]3s^2$ (D) $[Ar]3d^{10} 4s^2 4p^6$
8. Which element in Period 5 of the Periodic Table is a transition element?
(A) Sr (C) Ag
(B) Sb (D) Xe
9. Which of the following atoms has the largest atomic radius?
(A) Na (C) Mg
(B) K (D) Ca
10. Which noble gas has the highest first ionization energy?
(A) radon (C) neon
(B) krypton (D) helium

11. Which sequence of elements is arranged in order of decreasing atomic radii?
(A) Al, Si, P (C) Cl, Br, I
(B) Li, Na, K (D) N, C, B
12. Which list of elements from Group 2 on the Periodic Table is arranged in order of increasing atomic radius?
(A) Be, Mg, Ca (C) Ba, Ra, Sr
(B) Ca, Mg, Be (D) Sr, Ra, Ba
13. As each successive element in Group 15 of the Periodic Table is considered in order of increasing atomic number, the atomic radius
(A) decreases (C) remains the same
(B) increases (D) increases
14. The strength of an atom's attraction for the electrons in a chemical bond is the atom's
(A) electronegativity (C) heat of reaction
(B) ionization energy (D) heat of formation
15. Which properties are most common in nonmetals?
(A) low ionization energy and low electronegativity
(B) low ionization energy and high electronegativity
(C) high ionization energy and low electronegativity
(D) high ionization energy and high electronegativity
16. Which Group 17 element has the least attraction for electrons?
(A) F (C) Br
(B) Cl (D) I
17. Which element in Group 16 has the greatest tendency to gain electrons?
(A) Te (C) S
(B) Se (D) O
18. The Group 17 element with the highest electronegativity is
(A) fluorine (C) bromine
(B) chlorine (D) iodine
19. As the elements of Group 1 on the Periodic Table are considered in order of increasing atomic radius, the ionization energy of each successive element generally
(A) decreases (C) remains the same
(B) increases (D) increases
20. The amount of energy required to remove the outermost electron from a gaseous atom in the ground state is known as
(A) first ionization energy (C) conductivity
(B) activation energy (D) electronegativity
21. Which element is a member of the halogen family?
(A) K (C) I
(B) B (D) S

Periodic Table Trends Answer Honors

Steven S. Zumdahl



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Table of Contents Periodic Table Trends Answer Honors

1. Understanding the eBook Periodic Table Trends Answer Honors
 - The Rise of Digital Reading Periodic Table Trends Answer Honors
 - Advantages of eBooks Over Traditional Books
2. Identifying Periodic Table Trends Answer Honors
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Periodic Table Trends Answer Honors
 - User-Friendly Interface
4. Exploring eBook Recommendations from Periodic Table Trends Answer Honors
 - Personalized Recommendations
 - Periodic Table Trends Answer Honors User Reviews and Ratings
 - Periodic Table Trends Answer Honors and Bestseller Lists
5. Accessing Periodic Table Trends Answer Honors Free and Paid eBooks
 - Periodic Table Trends Answer Honors Public Domain eBooks
 - Periodic Table Trends Answer Honors eBook Subscription Services
 - Periodic Table Trends Answer Honors Budget-Friendly Options

6. Navigating Periodic Table Trends Answer Honors eBook Formats
 - ePub, PDF, MOBI, and More
 - Periodic Table Trends Answer Honors Compatibility with Devices
 - Periodic Table Trends Answer Honors Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Periodic Table Trends Answer Honors
 - Highlighting and Note-Taking Periodic Table Trends Answer Honors
 - Interactive Elements Periodic Table Trends Answer Honors
8. Staying Engaged with Periodic Table Trends Answer Honors
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Periodic Table Trends Answer Honors
9. Balancing eBooks and Physical Books Periodic Table Trends Answer Honors
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Periodic Table Trends Answer Honors
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Periodic Table Trends Answer Honors
 - Setting Reading Goals Periodic Table Trends Answer Honors
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Periodic Table Trends Answer Honors
 - Fact-Checking eBook Content of Periodic Table Trends Answer Honors
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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