

## Chapter 1: Limiting Reagents and Percent Yield (Pages 1-4)

**Remember:** Limiting reagent is the reactant that will completely run out first in a chemical reaction.

**Excess:** All reagents except the limiting reagent will be present in excess. Excess is the amount of reagent that is not used in the reaction.

Remember: Limiting reagent is the reactant that will completely run out first in a chemical reaction. Excess is the amount of reagent that is not used in the reaction.



The reaction is balanced. The coefficients are 2, 1, and 2.

2H <sub>2</sub>	2 x 2.016 g/mol
O <sub>2</sub>	2 x 31.998 g/mol
2H <sub>2</sub> O	2 x 18.015 g/mol
2H <sub>2</sub>	2 x 2.016 g/mol

1. 2.016 g/mol is the molar mass of H<sub>2</sub>. 2.016 g/mol is the molar mass of H<sub>2</sub>. 2.016 g/mol is the molar mass of H<sub>2</sub>. 2.016 g/mol is the molar mass of H<sub>2</sub>.



2. 31.998 g/mol is the molar mass of O<sub>2</sub>. 31.998 g/mol is the molar mass of O<sub>2</sub>. 31.998 g/mol is the molar mass of O<sub>2</sub>. 31.998 g/mol is the molar mass of O<sub>2</sub>.

3. 18.015 g/mol is the molar mass of H<sub>2</sub>O. 18.015 g/mol is the molar mass of H<sub>2</sub>O. 18.015 g/mol is the molar mass of H<sub>2</sub>O. 18.015 g/mol is the molar mass of H<sub>2</sub>O.



4. 18.015 g/mol is the molar mass of H<sub>2</sub>O. 18.015 g/mol is the molar mass of H<sub>2</sub>O. 18.015 g/mol is the molar mass of H<sub>2</sub>O. 18.015 g/mol is the molar mass of H<sub>2</sub>O.



# Percent Yield And Limiting Reactants Quiz

**Jessica J Manson**



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