

$$M = \frac{\# \text{ of moles}}{V (\text{in L})}$$

Molar Concentration

Part A: Describing Molar Concentration

1. Sodium chloride was dissolved in water to produce a 1.5M solution. Explain what this concentration tells us about the NaCl solution. How might a chemist use this ratio?

$$\frac{1.5 \text{ mol of NaCl}}{1 \text{ L of water}}$$

if you need a certain amount of NaCl you can take a certain volume of the solution.

2. A 45.3 g sample of potassium nitrate is dissolved in enough water to make 225 mL of solution. Determine the molar concentration of the potassium nitrate.

$$\begin{aligned} \text{KNO}_3 : \text{m.m.} &= 101.11 \text{ g/mol} & 0.45 \text{ mol KNO}_3 \\ V &= 0.225 \text{ L} & \frac{0.45 \text{ mol KNO}_3}{0.225 \text{ L}} = 2 \text{ M} \end{aligned}$$

3. Find the molarity of a solution made from 275 g of CuSO_4 dissolved in enough water to make 4.25 L.

$$\begin{aligned} 275 \text{ g CuSO}_4 & \quad \frac{275 \text{ g}}{159.65} = 1.72 \text{ mol CuSO}_4 \\ \text{m.m.} & \quad \frac{1.72 \text{ mol}}{4.25 \text{ L}} = 0.40 \text{ M CuSO}_4 \\ 63.55 & \\ + 32.1 & \\ \hline 95.65 & \end{aligned}$$

4. An alcoholic iodine solution ("tincture" of iodine) is prepared by dissolving 5.15 g of iodine crystals in enough alcohol to make a volume of 225 mL. Calculate the molarity of iodine in the solution.

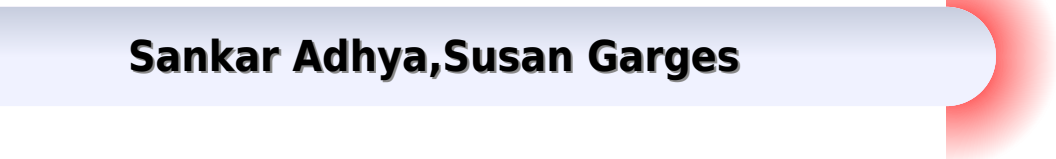
$$\begin{aligned} 5.15 \text{ g I}_2 & \quad \frac{0.02 \text{ mol I}_2}{0.225 \text{ L}} = 0.089 \text{ M I}_2 \\ 253.8 & \quad \frac{5.15 \text{ g I}_2}{253.8} = 0.02 \text{ mol I}_2 \end{aligned}$$

5. What final volume would be needed in order to prepare a 0.25 M NaCl solution from 5.2 g of NaCl (s)?

$$\begin{aligned} \text{m.m. NaCl} : & \quad \frac{5.2 \text{ g}}{58.44 \text{ g/mol}} = 0.089 \text{ mol NaCl} & M = \frac{\# \text{ of mol}}{V (\text{in L})} \\ 22.99 & & V (\text{in L}) = \frac{\# \text{ of mol}}{M} \\ 35.45 & & \\ \hline 58.44 & & \frac{0.089 \text{ mol}}{0.25 \text{ M}} = \boxed{0.36 \text{ L}} \end{aligned}$$

Part A Describing Molar Concentration

Sankar Adhya, Susan Garges



Part A Describing Molar Concentration:

Laboratory Manual for Principles of General Chemistry J. A. Beran, Mark Lassiter, 2022-08-16 Laboratory Manual for Principles of General Chemistry 11th Edition covers two semesters of a general chemistry laboratory program The material focuses on the lab experiences that reinforce the concepts that not all experimental conclusions are the same and depend on identifying an appropriate experimental procedure selecting the proper apparatus employing the proper techniques systematically analyzing and interpreting the data and minimizing inherent variables As a result of good data a scientific and analytical conclusion is made which may or may not be right but is certainly consistent with the data Experiments write textbooks textbooks don't write experiments A student's scientific literacy grows when experiences and observations associated with the scientific method are encountered Further experimentation provides additional cause effect observations leading to an even better understanding of the experiment The 11th edition's experiments are informative and challenging while offering a solid foundation for technique safety and experimental procedure The reporting and analysis of the data and the pre and post lab questions focus on the intuitiveness of the experiment The experiments may accompany any general chemistry textbook and are compiled at the beginning of each curricular unit An Additional Notes column is included in each experiment's Report Sheet to provide a space for recording observations and data during the experiment Continued emphasis on handling data is supported by the Data Analysis section

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Table of Contents Part A Describing Molar Concentration

1. Understanding the eBook Part A Describing Molar Concentration
 - The Rise of Digital Reading Part A Describing Molar Concentration
 - Advantages of eBooks Over Traditional Books
2. Identifying Part A Describing Molar Concentration
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Part A Describing Molar Concentration
 - User-Friendly Interface
4. Exploring eBook Recommendations from Part A Describing Molar Concentration
 - Personalized Recommendations
 - Part A Describing Molar Concentration User Reviews and Ratings
 - Part A Describing Molar Concentration and Bestseller Lists
5. Accessing Part A Describing Molar Concentration Free and Paid eBooks
 - Part A Describing Molar Concentration Public Domain eBooks
 - Part A Describing Molar Concentration eBook Subscription Services
 - Part A Describing Molar Concentration Budget-Friendly Options
6. Navigating Part A Describing Molar Concentration eBook Formats
 - ePub, PDF, MOBI, and More
 - Part A Describing Molar Concentration Compatibility with Devices
 - Part A Describing Molar Concentration Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Part A Describing Molar Concentration
 - Highlighting and Note-Taking Part A Describing Molar Concentration
 - Interactive Elements Part A Describing Molar Concentration
8. Staying Engaged with Part A Describing Molar Concentration

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Part A Describing Molar Concentration
- 9. Balancing eBooks and Physical Books Part A Describing Molar Concentration
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Part A Describing Molar Concentration
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Part A Describing Molar Concentration
 - Setting Reading Goals Part A Describing Molar Concentration
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Part A Describing Molar Concentration
 - Fact-Checking eBook Content of Part A Describing Molar Concentration
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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