QUESTION 5

A learner wants to determine the molecular formula of a compound found in cigarettes. The learner asks for your help with the calculation. The following is known about the compound: It is responsible for cigarettes being such an addictive substance, it has a molecular mass of 162,2 g-mol⁻¹ and its percentage composition is 74,07% C: 8,65% H and 17,28% N.

- 5.1 Name the compound being referred to here. (2)
- 5.2 As much as the substance referred to in QUESTION 5.1 is responsible for the addiction of smokers to cigarettes, it is used in some products to help them overcome their addiction. Name ONE example of such a product. (2)
- 5.3 Cigarette packs contain various warnings on them concerning the harmful effects of smoking. Name ONE harmful effect of smoking on humans. (1)
- 5.4 Show, by means of calculation, how you would go about helping the learner determine the molecular formula of this compound. (7) [12]

QUESTION 6

Many household substances are either acids or bases. Some examples are: vinegar, baking soda (bicarbonate of soda), caustic soda, lemon juice.

- 6.1 Use the examples provided to name ONE substance which is an acid and give a characteristic of it to support your answer. (3)
- 6.2 Give the formula for baking soda (bicarbonate of soda). (2)
- 6.3 A group of learners decide to do a titration to find the concentration of a sample of vinegar. To do the titration, they prepare a 250 cm³ standard solution of NaOH of concentration 0,2 mol-dm⁻³. It is found that during the titration 8,5 cm³ of NaOH reacted with 4 cm³ of the vinegar sample. The balanced chemical equation for the reaction is:

CH₃COOH + NaOH → CH₃COONa + H₂O

- 6.3.1 Give an investigative question for this investigation. (2)
- 6.3.2 What is the scientific name for vinegar? (2)
- 6.3.3 Calculate the concentration of the vinegar sample. (6)

(3)

- 6.4 A certain antacid tablet contains MgCO₃. When taken, the antacid tablet neutralises the excess hydrochloric acid in the stomach.
 - 6.4.1 Write a balanced chemical equation for the reaction between MgCO₃ and the hydrochloric acid.

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Michael A. Hitt, C. Chet Miller, Adrienne Colella, Maria Triana

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