

Answer all questions on the paper provided. You may use a non-programmable calculator & refer to the data sheet where necessary.

**QUESTION 1**

There are four possible options for each answer in the following questions. Each question has only ONE correct answer. Choose the correct answer and write only A, B, C or D next to the question number.

- 1.1 Which ONE of the following statements is true of a body that is accelerating at a constant rate?
- A The rate of change of velocity with time remains constant.
  - B The velocity of the body increases by increasing amounts in equal time intervals.
  - C The position changes by the same amount in equal time intervals.
  - D The rate of change of position with time remains constant.
- (2)

- 1.2 Joseph is leaning on the railing of a tall building. His cellphone drops out of his hands and falls to the ground. Which of the following combinations of velocity and acceleration for the motion of the cellphone is correct?

	Velocity	Acceleration
A	Increases	Increases
B	Decreases	Increases
C	Increases	Remains constant
D	Remains constant	Remains constant

(2)

- 1.3 The distance between two consecutive crests in a wave train is 8 cm. If two complete waves pass a point in 1 s, then the velocity of the wave is:
- A  $16 \text{ cm s}^{-1}$
  - B  $8 \text{ cm s}^{-1}$
  - C  $4 \text{ cm s}^{-1}$
  - D  $1,25 \text{ cm s}^{-1}$
- (2)

# Physical Sciences Control Test No

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## **Table of Contents Physical Sciences Control Test No**

1. Understanding the eBook Physical Sciences Control Test No
  - The Rise of Digital Reading Physical Sciences Control Test No
  - Advantages of eBooks Over Traditional Books
2. Identifying Physical Sciences Control Test No
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Physical Sciences Control Test No
  - User-Friendly Interface
4. Exploring eBook Recommendations from Physical Sciences Control Test No
  - Personalized Recommendations
  - Physical Sciences Control Test No User Reviews and Ratings
  - Physical Sciences Control Test No and Bestseller Lists
5. Accessing Physical Sciences Control Test No Free and Paid eBooks
  - Physical Sciences Control Test No Public Domain eBooks
  - Physical Sciences Control Test No eBook Subscription Services

- Physical Sciences Control Test No Budget-Friendly Options
- 6. Navigating Physical Sciences Control Test No eBook Formats
  - ePub, PDF, MOBI, and More
  - Physical Sciences Control Test No Compatibility with Devices
  - Physical Sciences Control Test No Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Physical Sciences Control Test No
  - Highlighting and Note-Taking Physical Sciences Control Test No
  - Interactive Elements Physical Sciences Control Test No
- 8. Staying Engaged with Physical Sciences Control Test No
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Physical Sciences Control Test No
- 9. Balancing eBooks and Physical Books Physical Sciences Control Test No
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Physical Sciences Control Test No
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Physical Sciences Control Test No
  - Setting Reading Goals Physical Sciences Control Test No
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Physical Sciences Control Test No
  - Fact-Checking eBook Content of Physical Sciences Control Test No
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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 $\text{H}_2\text{O}(\text{g}) \rightarrow \text{H}_2(\text{g}) + \frac{1}{2} \text{O}_2(\text{g}) \quad \Delta h = 71.2 \text{ kJ}$   
 $\text{C}_2\text{H}_6(\text{g}) + \frac{7}{2} \text{O}_2(\text{g}) \rightarrow 2 \text{ CO}_2(\text{g}) + 3 \text{ H}_2\text{O}(\text{g}) \quad \Delta h = 283 \text{ kJ}$

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 $\text{C}_2\text{H}_4(\text{g}) + 3 \text{ O}_2(\text{g}) \rightarrow 2 \text{ CO}_2(\text{g}) + 2 \text{ H}_2\text{O}(\text{l}) \quad \Delta h = 1411 \text{ kJ/mole}$   
 $\text{C}_2\text{H}_6(\text{g}) + 7/2 \text{ O}_2(\text{g}) \rightarrow 2 \text{ CO}_2(\text{g}) + 3 \text{ H}_2\text{O}(\text{l}) \quad \Delta h = 1560 \text{ kJ/mole}$   
 $\text{H}_2(\text{g}) + 1/2 \text{ O}_2(\text{g}) \rightarrow \text{H}_2\text{O}(\text{l}) \quad \Delta h = 285 \text{ kJ/mole}$

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values to determine the heat of reaction for the reaction  $\text{C}_2\text{H}_4$

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web thus for reaction 9 2 in which 1 mole of  $\text{H}_2\text{O}$  and 1 mole of  $\text{OH}^-$  combine to form 2 moles of  $\text{H}_2\text{O}$   $\Delta H = -55.8 \text{ kJ}$  as shown note that the enthalpy change  $\Delta H_{\text{products}} - \Delta H_{\text{reactants}}$  is positive if heat is absorbed that is if  $\Delta H_{\text{products}} > \Delta H_{\text{reactants}}$

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