Honors Chemistry - Periodic Trends Worksheet

Name: _____

1. Circle the element with the largest atomic radius and put a square around the element with the smallest atomic radius:

    Cu   (K)   Ni   [Br]

    a. Explain why you made these choices: **All of the elements are in the same period. The trend in atomic radius as you go across a period is DECREASING. Therefore, the element on the far left (K) is the largest, and the element on the far right (Br) is the smallest.**

2. Circle the element with the highest ionization energy and put a square around the element with the lowest ionization energy:

    Cu   [K]   Ni   (Br)

    a. Explain why you made these choices: **All of the elements are in the same period. The trend in ionization energy as you go across a period is INCREASING. Therefore, the element on the far left (K) has the lowest ionization energy, and the element on the far right (Br) has the highest ionization energy.**

3. Circle the element with the highest electronegativity and put a square around the element with the lowest electronegativity:

    Cu   [K]   Ni   (Br)

    a. Explain why you made these choices: **All of the elements are in the same period. The trend in electronegativity as you go across a period is INCREASING. Therefore, the element on the far left (K) has the lowest electronegativity, and the element on the far right (Br) has the highest electronegativity.**

4. For each of the following groups: Circle the element with the largest atomic radius and put a square around the element with the smallest atomic radius:

5.

    a. O   C   (Be)   [Ne]   **Same Period**

    b. Na   Rb   (Fr)   [H]   **Same Group**

    c. (Pb)   [C]   Sn   Si   **Same Group**

    d. Au   W   S   (Fr)   [Ne]   Zn   **Challenge**

6. For each of the following groups: Circle the element with the highest ionization energy and put a square around the element with the lowest ionization energy:

    a. O   C   [Be]   (Ne)   **Same Period**

    b. Na   Rb   [Fr]   (H)   **Same Group**

    c. [Pb]   (C)   Sn   Si   **Same Group**

    d. Au   W   S   [Fr]   (Ne)   Zn   **Challenge**

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