

The Periodic Law

SECTION 2

Select the correct answer for the following questions for the space provided.

1. _____ for the modern periodic table, elements are ordered
 - (a) according to increasing atomic mass
 - (b) according to Mendeleev's original design
 - (c) according to increasing atomic number
 - (d) based on how they were discovered
2. _____ Mendeleev's original short periodic table gave useful chemical properties of elements appeared in regular intervals and elements discovered later completed intervals of his design.
 - (a) slowly
 - (b) not long
 - (c) almost instantly
 - (d) almost never
3. _____ The modern periodic law states that
 - (a) all new elements will be added soon and be listed in the table they were given
 - (b) the physical and chemical properties of elements are functions of atomic number
 - (c) elements exhibit properties of both periods and groups
 - (d) the chemical properties of elements can be grouped according to periodicity but physical properties cannot
4. _____ The discovery of the noble gases changed Mendeleev's periodic table.
 - (a) adding a row
 - (b) adding a column
 - (c) adding a new section
 - (d) adding a group
 - (e) adding a period
5. _____ The noble gases that property of the noble gases is that they are
 - (a) unstable
 - (b) inert
 - (c) radioactive
 - (d) highly reactive
6. _____ Likewise, the first column in Group 1 has characteristic elements of 1. The second column in the group has characteristic elements of
 - (a) 2
 - (b) 10
 - (c) 18
 - (d) 8
7. How groups in Mendeleev's periodic table of 18 and the elements number of 50
 - (a) 10
 - (b) 18
 - (c) 20
 - (d) 22
8. How many elements are in the group?
 - (a) 10
 - (b) 18
 - (c) 20
 - (d) 22
9. How many elements are in the group?
 - (a) 10
 - (b) 18
 - (c) 20
 - (d) 22
10. How many elements are in the group?
 - (a) 10
 - (b) 18
 - (c) 20
 - (d) 22

Periodic Law Chapter 5 Test Answers

Janet O'Sullivan, Jonathan Hilliard



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