#### Trends on the Periodic Table

#### Part 1: Identifying chemical properties/characteristics of the elements

For each of the following elements, draw a caricature or cartoon above the description to show its "chemical personality" based on the description provided. (Note: See the Neon example for ideas.)



# Atomic Rodii Size:

1. Explain atomic radius: A measure of the size (volume) of an atom. Determined by measuring half of the distance between two identical atoms bonded together.

2. Using the information in figure 5-13 (pg 141) in your book, write the atomic radii size for alkali metals & period #2 on the periodic table to your right.



- 3. What happens to atomic radii as you move down a column and across a row? On the arrows on the diagram, indicate the trend. (Options: gets bigger OR gets smaller)
- 4. Of all the elements on the periodic table, identify & circle the largest & smallest in size.
- Rank the following elements from largest to smallest in atomic radii: a) Be, Ba, Mg = Ba, Mg, Be
- 6. How can you explain the trend:

a. Why do they get smaller as you go across a period? Additional protons in the nucleus "pull" on

in creases the attraction of the nucleus and therefore cloud close to the nucleus resulting

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Scott C. Dulebohn

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#### **Table of Contents Periodic Trends Worksheet Pogil Extension Questions**

- 1. Understanding the eBook Periodic Trends Worksheet Pogil Extension Questions
  - The Rise of Digital Reading Periodic Trends Worksheet Pogil Extension Questions
  - Advantages of eBooks Over Traditional Books
- 2. Identifying Periodic Trends Worksheet Pogil Extension Questions
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Periodic Trends Worksheet Pogil Extension Questions
  - User-Friendly Interface
- 4. Exploring eBook Recommendations from Periodic Trends Worksheet Pogil Extension Questions
  - Personalized Recommendations
  - Periodic Trends Worksheet Pogil Extension Questions User Reviews and Ratings
  - Periodic Trends Worksheet Pogil Extension Questions and Bestseller Lists
- 5. Accessing Periodic Trends Worksheet Pogil Extension Questions Free and Paid eBooks

- o Periodic Trends Worksheet Pogil Extension Questions Public Domain eBooks
- Periodic Trends Worksheet Pogil Extension Questions eBook Subscription Services
- Periodic Trends Worksheet Pogil Extension Questions Budget-Friendly Options
- 6. Navigating Periodic Trends Worksheet Pogil Extension Questions eBook Formats
  - o ePub, PDF, MOBI, and More
  - Periodic Trends Worksheet Pogil Extension Questions Compatibility with Devices
  - Periodic Trends Worksheet Pogil Extension Questions Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Periodic Trends Worksheet Pogil Extension Questions
  - Highlighting and Note-Taking Periodic Trends Worksheet Pogil Extension Questions
  - Interactive Elements Periodic Trends Worksheet Pogil Extension Questions
- 8. Staying Engaged with Periodic Trends Worksheet Pogil Extension Questions
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Periodic Trends Worksheet Pogil Extension Questions
- 9. Balancing eBooks and Physical Books Periodic Trends Worksheet Pogil Extension Questions
  - Benefits of a Digital Library
  - o Creating a Diverse Reading Collection Periodic Trends Worksheet Pogil Extension Questions
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Periodic Trends Worksheet Pogil Extension Questions
  - Setting Reading Goals Periodic Trends Worksheet Pogil Extension Questions
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Periodic Trends Worksheet Pogil Extension Questions
  - Fact-Checking eBook Content of Periodic Trends Worksheet Pogil Extension Questions
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development

- Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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