

Annotated periodic table of the elements

1	H 1.01 Hydrogen		2	Group name* 1 Hydrogen & the alkali metals 2 Alkaline earth metals 11 Coinage metals 12 Volatile metals 13 Icosagens										14 Crystallogens 15 Pnictogens 16 Chalcogens 17 Halogens 18 Helium group	13	B 10.8 Boron	14	C 12.0 Carbon	15	N 14.0 Nitrogen	16	O 16.0 Oxygen	17	F 19.0 Fluorine	He 4.00 Helium		
2	Li 6.9 Lithium	Be 9.0 Beryllium	Most common oxidation states: electron affinity (kJ/mol-Richter)																								
3	Na 22.9 Sodium	Mg 24.3 Magnesium																									
4	K 39.1 Potassium	Ca 40.1 Calcium	*Groups 3 to 10 are named after their first members i.e. Group 3 is the Scandium group																								
5	Rb 85.5 Rubidium	Sr 87.6 Strontium	Stable number and symbol Stable weight and name																								
6	Cs 133 Cesium	Ba 137 Barium	57 to 70																								
7	Fr 223 Francium	Ra 226 Radium	89 to 102																								

CLASS COUNTERPARTS		
		Nonmetals
	Pre-transition	Corrosive
	Alkali earth	
	Alkaline	
Transition		Intermediate
Frontier		Metalloid
Noble (Ru-Pt, Os-Au)		Noble gas

57 La 138.91 Lanthanum	58 Ce 140.12 Cerium	59 Pr 140.91 Praseodymium	60 Nd 144.24 Neodymium	61 Pm 144.91 Promethium	62 Sm 150.36 Samarium	63 Eu 151.96 Europium	64 Gd 157.25 Gadolinium	65 Tb 158.93 Terbium	66 Dy 162.50 Dysprosium	67 Ho 164.93 Holmium	68 Er 167.26 Erbium	69 Tm 168.93 Thulium	70 Yb 173.05 Ytterbium
89 Ac 227.03 Actinium	90 Th 232.04 Thorium	91 Pa 231.04 Protactinium	92 U 238.03 Uranium	93 Np 237.05 Neptunium	94 Pu 244.06 Plutonium	95 Am 243.06 Americium	96 Cm 247.07 Curium	97 Bk 247.07 Berkelium	98 Cf 251.08 Californium	99 Es 252.08 Einsteinium	100 Fm 257.10 Fermium	101 Md 258.11 Mendelevium	102 Ds 261.10 Darmstadtium

Active metals

Pre-transition
Light, reactive, mostly of low mechanical strength, melting and boiling points. The group 1 members are easily cut with a knife. Only Be, Mg and Al have any structural uses. Most form basic oxides (Be and Al are amphoteric).

Rare earths
Closely resemble Ca, Sr and Ba, but many are heavier and they form mostly pale coloured compounds. Similar to one another, esp. La to Lu (aka the lanthanides), and hard to separate for that reason. Abundant but widely dispersed; commercially viable concentrations are rare. Oxides are strongly to moderately basic. Sc, Yb and Lu exhibit weak amphotericism.

Actinides
Soft, dense and reactive metals. Those from Ra onwards are synthetic. Early actinides (Th-Am) show some similarities to transition metals and have basic, amphoteric, or acidic oxides; late actinides are more like lanthanides. Many actinides form coloured compounds.

Transition metals
Moderate to high density metals, with high melting and boiling points, many have high hardness and mechanical strength. Those from Rf onwards are synthetic. Chemically show variable valency and strong tendency to form coordination and colored compounds. Oxides are basic, amphoteric or acidic, depending on oxidation state.

Noble metals
Resistant to corrosion and oxidation, unlike base metals such as Fe, Ni, Cu and Zn. Commonly found together in native form. Good catalytic properties. Very weakly electropositive; insoluble oxides.

Frontier metals
Soft (or brittle), low strength, and with melting points lower than transition metals. Crystalline structures tend to show directional bonding, with generally greater complexity or lower nearest neighbours than other metals. All used or tried as soft metal lubricants. Most are chemically weak; a minority are moderately electropositive. Characterised, to varying degrees, by covalent bonding tendencies, amphotericism, and anionic species such as zincates, stannates, and bismuthates.

Metalloids
Metallic looking brittle, fair electrical conductors. All known in semiconducting forms. Moderate ionisation energies (750-1,000 kJ/mol), and electronegativity values (1.8-2.3). Mostly behave chemically as non-metals. Can form alloys with metals. Amphoteric or weakly acidic oxides.

Intermediate nonmetals
H and N are colourless gases; C, P, and Se are metalloids looking solids while S is yellow. The latter four are brittle but all known in malleable, plastic or ductile forms. Moderate to high ionisation energies, low to high electron affinities, moderate to high electronegativity values, relatively poor to moderately strong oxidising agents, tend to form covalent or metallic to near-metallic compounds with metals. Their oxides are acidic, or neutral (H₂O, CO, NO, N₂O).

** as liquid P, the most stable form in ambient conditions.*

Group 1
The alkali are Li-F, and the alkaline earths are La-Be, and Al-Li trends (i.e. Li-Li). Trends going down: Li-F and Al-Li trends (i.e. Li-Li). Trends going down: Li-F and Al-Li trends (i.e. Li-Li).

- Li trend, resemble those of groups 1-2
- Li trend, more resemble those of groups 4-10

Corrosive nonmetals
O, F, and Cl are, respectively, colourless, pale yellow, and pale yellow-green gases. Br is a red-brown liquid. I is a metallic looking solid. Chemically tend to have to high ionisation energies, electron affinities, and electronegativity values, and be relatively strong oxidising agents. Manifestations include oxygen's major association with the ubiquitous processes of corrosion and combustion, and the intrinsically corrosive nature of the non-metallic halogens. All five tend to form predominantly ionic compounds with metals. Their oxides are acidic.

Noble gases
Colourless, odourless, non-flammable gases with very low chemical reactivity. The first noble gas compound, of approximate composition XePF₆, was not prepared until 1962. More than half-a-thousand noble gas compounds were synthesised in the ensuing half-century. Compounds of He, Ne, and Ar, in ambient conditions, are not known.

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**John Krenos, Joseph Potenza, Carl
Hoeger**



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