Periodicity Chemistry Worksheet

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1. Which are metals? Circle your answers: C. Na. F. Cs. Ba, Ni.

Which metal in the list above has the most metallic character? Explain.

Cs - It is the metal with the † shielding & | ENC

Write the charge that each of the following atoms will have when it has a complete set of valence electrons forming an ion.

O 4 Na 1 F 1 N 3 Ca 2 Ar 6

What is the most common oxidation number for calcium? Explain.
 Ca¹² – Calcium has 2 valence electrons, it will lose 2 electrons.

Name two more elements with that oxidation number and explain your choice.

Any of the two: Be, Mg, Se, Ba, Ra – Same valence electrons / same family

- 4. What element in period 3 is a metalloid? Silicon
- When element with atomic number 118 is discovered, what family will it be in? Noble Gas.

Use the following word bank for questions 6 through 13

(Obviously, you will need to use several words more than once! ○)
Alkali metals Alkaline earth metals Halogens
Noble gases Transition metals Noble gases

- The __Alkali metals_____ have a single electron in the highest energy level.
- The __Alkaline earth metals__ achieve the electron configurations of noble gases by losing two electrons.
- The __transition metals__ are metals that can hold up to 10 electrons in their sublevel shape
- The _halogens ____ achieve the electron configuration of noble gases by gaining one electron.
- The <u>noble gases</u> have full s and p orbitals in the highest occupied energy levels.
- 11. The __noble gases_____ are stable and un-reactive.
- 12. The __halogens _____ are highly reactive and readily form salts with metals.
- 13. The __Alkaline earth metals are metals that are more reactive than the transition elements but less reactive than the alkali metals.
- 14. Predict the oxidation number based on the electron configuration shown.

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M Carnoy

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