

Graphing Periodic Trends Activity

Name: _____ Period: _____

Use the information in these tables to complete the graph as described below

Symbol	Atomic Radius (Picometers)	First Ionization energy (kilojoules/mole)	Electronegativity (4-point scale)
H	31	1312	2.1
Li	128	520	1.0
Na	166	496	0.9
K	203	410	0.8
Rb	220	403	0.8
Cs	244	376	0.7

Symbol	Atomic Radius (Picometers)	First Ionization energy (kilojoules/mole)	Electronegativity (4-point scale)
Na	166	496	0.9
Mg	141	738	1.2
Al	121	578	1.5
Si	111	787	1.8
P	107	1012	2.1
S	105	1000	2.5
Cl	102	1251	3.0
Ar	106	1521	---

- Using colored pencils or pen, list under the symbol (in this order!) the Atomic Radius, First Ionization Energy, and Electronegativity.
- Use a different color for each property. Example: Write all of the Atomic radius values in red, all of the First ionization energies in green, and all of the Electronegativities in blue. Of course, you can pick any colors that are available, as long as you are consistent.
- Observe the trends in each property as you go down the Alkali metal group, and as you go across Period 3.
- Write out each of the statements written on the opposite side, completing each statement with the observed trend (increase or decrease)

H	Across a period (→) atomic radius tends to _____. Across a period (→) first ionization energy tends to _____. Across a period (→) electronegativity tends to _____.						
Li							
Na 166 496 0.9	Mg	Al	Si	P	S	Cl	Ar
K	Down a Group (↓) atomic radius tends to _____. Down a Group (↓) first ionization energy tends to _____. Down a Group (↓) electronegativity tends to _____.						
Rb							
Cs							

Periodic Trends Activity Packet Answers

J Dewey



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Table of Contents Periodic Trends Activity Packet Answers

1. Understanding the eBook Periodic Trends Activity Packet Answers
 - The Rise of Digital Reading Periodic Trends Activity Packet Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Periodic Trends Activity Packet Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Periodic Trends Activity Packet Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Periodic Trends Activity Packet Answers
 - Personalized Recommendations
 - Periodic Trends Activity Packet Answers User Reviews and Ratings
 - Periodic Trends Activity Packet Answers and Bestseller Lists
5. Accessing Periodic Trends Activity Packet Answers Free and Paid eBooks

- Periodic Trends Activity Packet Answers Public Domain eBooks
 - Periodic Trends Activity Packet Answers eBook Subscription Services
 - Periodic Trends Activity Packet Answers Budget-Friendly Options
6. Navigating Periodic Trends Activity Packet Answers eBook Formats
- ePub, PDF, MOBI, and More
 - Periodic Trends Activity Packet Answers Compatibility with Devices
 - Periodic Trends Activity Packet Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Periodic Trends Activity Packet Answers
 - Highlighting and Note-Taking Periodic Trends Activity Packet Answers
 - Interactive Elements Periodic Trends Activity Packet Answers
8. Staying Engaged with Periodic Trends Activity Packet Answers
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Periodic Trends Activity Packet Answers
9. Balancing eBooks and Physical Books Periodic Trends Activity Packet Answers
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Periodic Trends Activity Packet Answers
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Periodic Trends Activity Packet Answers
- Setting Reading Goals Periodic Trends Activity Packet Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Periodic Trends Activity Packet Answers
- Fact-Checking eBook Content of Periodic Trends Activity Packet Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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