

Concepts in Chemistry

The Periodic Law

SECTION 2

Select the correct answer for the following questions for the space provided.

1. _____ for the modern periodic table, elements are ordered
 - (a) according to increasing atomic mass
 - (b) according to Mendeleev's original design
 - (c) according to increasing atomic number
 - (d) based on how they were discovered
2. _____ Mendeleev's original short periodic table gave better chemical properties of elements appeared in regular intervals whereas elements were arranged according to increasing
 - (a) density
 - (b) atomic number
 - (c) atomic mass
 - (d) electronegativity
3. _____ The modern periodic law states that
 - (a) all new elements with the same properties are listed in the same group
 - (b) the physical and chemical properties of elements are functions of their atomic number
 - (c) elements exhibit properties of both periods and groups
 - (d) the chemical properties of elements can be grouped according to periodicity but physical properties cannot
4. _____ The discovery of the noble gases changed Mendeleev's periodic table by adding a new
 - (a) period
 - (b) group
 - (c) series
 - (d) block
5. _____ The main reason the properties of the noble gases is that they are
 - (a) inert
 - (b) unreactive
 - (c) large gaseous molecules
 - (d) monatomic
6. _____ Likewise, the first element in Group 1, has an atomic number of 3. The second element in this group has an atomic number of
 - (a) 4
 - (b) 11
 - (c) 20
 - (d) 35
7. How groups in Mendeleev's periodic table of 18 and the atomic number of 55
 - (a) 11. How many periods are in this series?
 - (b) 11. How many elements are in this group?
 - (c) 11. How many elements are in this group?
 - (d) 11. How many elements are in this group?

Periodic Law Chapter 5 Test Answers

L Darling-Hammond



Periodic Law Chapter 5 Test Answers:

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