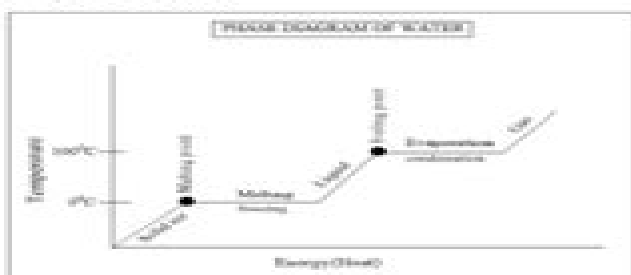


Topic 1: Stoichiometry

1.1 Introduction to Matter & Chemical Change

- **matter:** has mass, consumes volume, made up of particles in constant motion
- **atomic theory:** all matter is composed of atoms that cannot be created nor destroyed
- **states of matter:**
 - **solid:** fixed volume/shape; cannot be compressed; attractive forces hold particles in close-packed arrangement; particles vibrate in fixed positions
 - **liquid:** fixed volume, no fixed shape; cannot be compressed; weaker forces; particles move around
 - **gas:** no fixed volume/shape; can be compressed
- as temperature ↑, KE_{ave} ↑
 - Kelvin (K): SI unit of temperature; $^{\circ}\text{C} + 273$
- **changes of state:**



- **elements:** atoms of one type → form compounds
- **pure substance:** matter with constant composition → proportions are distinct
 - combine to form mixtures
 - homogeneous mixture: uniform composition and proportions
 - heterogeneous mixture: non-uniform composition and varying proportions
- **ion:** charged species
 - **anions:** negatively charged (contain excess electrons)
 - **cations:** positively charged (contains fewer electrons)
- **types of reactions:**
 - combination/synthesis rxns: combination of multiple substances
 - decomposition: breaking down of a single reactant
 - single displacement: one element replaces another in a compound
 - double displacement: swapping of elements

1.2 The Mole Concept

- **mole:** number of particles in a substance
 - fixed amount → Avogadro's number: 6.02×10^{23}
- **isotopes:** different number of neutrons/mass number; same number of protons
 - relative abundance: measure of the percentage of isotopes in a sample
 - relative atomic mass (A_r): weighted average of atomic masses of isotopes/relative abundance
 - relative molecular mass (M_r): combination of all A_r s
- **molar mass:** mass of 1 mol of substance (g/mol)

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1. Understanding the eBook Past Exam Papers Ib Chemistry
 - The Rise of Digital Reading Past Exam Papers Ib Chemistry
 - Advantages of eBooks Over Traditional Books
2. Identifying Past Exam Papers Ib Chemistry
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Past Exam Papers Ib Chemistry
 - User-Friendly Interface
4. Exploring eBook Recommendations from Past Exam Papers Ib Chemistry

- Personalized Recommendations
- Past Exam Papers Ib Chemistry User Reviews and Ratings
- Past Exam Papers Ib Chemistry and Bestseller Lists
- 5. Accessing Past Exam Papers Ib Chemistry Free and Paid eBooks
 - Past Exam Papers Ib Chemistry Public Domain eBooks
 - Past Exam Papers Ib Chemistry eBook Subscription Services
 - Past Exam Papers Ib Chemistry Budget-Friendly Options
- 6. Navigating Past Exam Papers Ib Chemistry eBook Formats
 - ePub, PDF, MOBI, and More
 - Past Exam Papers Ib Chemistry Compatibility with Devices
 - Past Exam Papers Ib Chemistry Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Past Exam Papers Ib Chemistry
 - Highlighting and Note-Taking Past Exam Papers Ib Chemistry
 - Interactive Elements Past Exam Papers Ib Chemistry
- 8. Staying Engaged with Past Exam Papers Ib Chemistry
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Past Exam Papers Ib Chemistry
- 9. Balancing eBooks and Physical Books Past Exam Papers Ib Chemistry
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Past Exam Papers Ib Chemistry
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Past Exam Papers Ib Chemistry
 - Setting Reading Goals Past Exam Papers Ib Chemistry
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Past Exam Papers Ib Chemistry

- Fact-Checking eBook Content of Past Exam Papers Ib Chemistry
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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