

- (b) A buffer is prepared by mixing 100 mL of 0.10 M acetic acid with 100 mL of 0.10 M sodium acetate.

- (c) The buffer is then a 100 mL solution of 0.10 M acetic acid and 0.10 M acetate.

The buffer is then a 100 mL solution of 0.10 M acetic acid and 0.10 M acetate.

In this problem, we should provide both of the general equations that would be used in including appropriate quantities and necessary calculations.

For the buffer solution with 0.10 M acetic acid and 0.10 M acetate.

$$[CH_3COOH] = 0.10 \text{ M} \quad [CH_3COO^-] = 0.10 \text{ M}$$

$$K_a = \frac{[CH_3COO^-][H^+]}{[CH_3COOH]} \quad [CH_3COO^-] = \frac{[CH_3COOH]}{K_a} = \frac{0.10 \text{ M}}{1.8 \times 10^{-5}} = 5.56 \times 10^{-3} \text{ M}$$

$$pH = -\log[H^+] = -\log(1.8 \times 10^{-5}) = 4.74$$

$$pH = 4.74$$

Reaction: $CH_3COOH + H_2O \rightleftharpoons CH_3COO^- + H_3O^+$
Initial: 0.10 M, 0.10 M, 0.00 M, 0.00 M
Change: -x, -x, +x, +x
Equilibrium: 0.10 - x, 0.10 - x, x, x
Since K_a is small, we can assume x is small compared to 0.10.
So, $0.10 - x \approx 0.10$
 $K_a = \frac{x \cdot x}{0.10 \cdot 0.10} = 1.8 \times 10^{-5}$
 $x^2 = 1.8 \times 10^{-5} \cdot 0.10 \cdot 0.10$
 $x = 1.34 \times 10^{-3} \text{ M}$
 $pH = -\log(1.34 \times 10^{-3}) = 2.87$

- (b) A buffer is prepared by mixing 100 mL of 0.10 M acetic acid with 100 mL of 0.10 M sodium acetate.

The buffer is then a 100 mL solution of 0.10 M acetic acid and 0.10 M acetate.



- (c) The buffer is then a 100 mL solution of 0.10 M acetic acid and 0.10 M acetate.

The buffer is then a 100 mL solution of 0.10 M acetic acid and 0.10 M acetate.

- (d) The buffer is then a 100 mL solution of 0.10 M acetic acid and 0.10 M acetate.

$$K_a = \frac{[CH_3COO^-][H^+]}{[CH_3COOH]} \quad [CH_3COO^-] = \frac{[CH_3COOH]}{K_a} = \frac{0.10 \text{ M}}{1.8 \times 10^{-5}} = 5.56 \times 10^{-3} \text{ M}$$

$$pH = -\log[H^+] = -\log(1.8 \times 10^{-5}) = 4.74$$

$$pH = 4.74$$

$$K_a = \frac{[CH_3COO^-][H^+]}{[CH_3COOH]} = 1.8 \times 10^{-5}$$

$$[CH_3COO^-] = \frac{[CH_3COOH]}{K_a} = \frac{0.10 \text{ M}}{1.8 \times 10^{-5}} = 5.56 \times 10^{-3} \text{ M}$$

$$pH = -\log[H^+] = -\log(1.8 \times 10^{-5}) = 4.74$$

$$pH = 4.74$$

$$K_a = \frac{[CH_3COO^-][H^+]}{[CH_3COOH]} = 1.8 \times 10^{-5}$$

$$[CH_3COO^-] = \frac{[CH_3COOH]}{K_a} = \frac{0.10 \text{ M}}{1.8 \times 10^{-5}} = 5.56 \times 10^{-3} \text{ M}$$

$$pH = -\log[H^+] = -\log(1.8 \times 10^{-5}) = 4.74$$

$$pH = 4.74$$

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