

**Instructions:-**

Question paper consist of 31 questions divided into four sections namely A, B, C, D.

- 1) **Section-A:** Q. No. 1 contains ten multiple choice types of questions carrying one mark each.  
Q.No.2 contains eight very short answer types of questions carrying one mark each.
- 2) **Section-B :** Q.No.3 to Q. No. 14 are Twelve questions carrying two mark each out of which any eight to be attempted.
- 3) **Section-C :** Q. No.15 to Q.No.26 are Twelve questions carrying three mark each out of which any eight are to be attempted.
- 4) **Section-D :** Q.No.27 to Q.No. 31 are Five questions carrying four mark each out of which any three to be attempted.

**Section- A**

**Q.1) Select and write the correct answer:**

**(10)**

- i. Which of the following compounds can NOT demonstrate the law of proportions?  
a) NO, NO<sub>2</sub>      b) CO, CO<sub>2</sub>      c) H<sub>2</sub>O, H<sub>2</sub>O<sub>2</sub>      d) Na<sub>2</sub>S, NaF
- ii. Which of the following temperature will read the same value on Celsius and Fahrenheit Scales  
a) -40      b) + 40      c) -80      d) -20
- iii. SI unit of the quantity electric current is .....  
a) Volts      b) Ampere      c) Candela      d) Newton
- iv. The no. of molecules in 22.4 cm<sup>3</sup> of Nitrogen Gas at STP is  
a)  $6.022 \times 10^{20}$       b)  $6.022 \times 10^{23}$       c)  $22.4 \times 10^{20}$       d)  $22.4 \times 10^{23}$
- v. The homogeneous series of alcohol has general molecular formula ?  
a) C<sub>n</sub> H<sub>2</sub>n+2OH      b) C<sub>n</sub> H<sub>2</sub>n OH      c) C<sub>n</sub> H<sub>2</sub>n+1OH      d) none
- vi. Find the molar mass of CH<sub>3</sub>CHO.  
a) 44      b) 40      c) 80      d) 30
- vii. Find the molar mass of CH<sub>3</sub>NO<sub>2</sub>.  
a) 61      b) 77      c) 98      d) 51
- viii. Atomic number of an elements is equal to  
a) P = e-      b) p = N      c) P = A      d) p = n = e

# Nmaths 2examination Paper

**RS Peters**



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