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**READING STREET**



# Reading Street 5th Grade Textbook

**Michal Rosen-Zvi**



## **Reading Street 5th Grade Textbook:**

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web limiting reactant theoretical and percentage yields key  $4\text{KO}_2 + 2\text{H}_2\text{O} \rightarrow 4\text{KOH} + 3\text{O}_2$  g if a reaction vessel contains 0 mol  $\text{KO}_2$  and 0 mol  $\text{H}_2\text{O}$  what is the limiting reactant

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the balanced equation shown below if the reaction of 90.6 grams of CO produces 36.7 grams of C<sub>3</sub>H<sub>8</sub> what is the percent yield  
 $3\text{CO} + 7\text{H}_2 \rightarrow \text{C}_3\text{H}_8 + 3\text{H}_2\text{O}$

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web start with 8.3 mole H<sub>2</sub>S and multiply it with 2 mole H<sub>2</sub>O over 2 mole H<sub>2</sub>S from the equation  
 $8.3 \text{ mole H}_2\text{S} \times \frac{2 \text{ mole H}_2\text{O}}{2 \text{ mole H}_2\text{S}} = 8.3 \text{ mole H}_2\text{O}$   
 8.3 mole H<sub>2</sub>S  $\times$  18g H<sub>2</sub>O / 1 mole H<sub>2</sub>O = 149.4g H<sub>2</sub>O is your theoretical yield so percent yield =  $\frac{137.1\text{g H}_2\text{O}}{149.4\text{g H}_2\text{O}} \times 100 = 91.77\%$   
 problems 6 & 8 are similar to problem 5

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web 1 csun.edu/hcchm001/introchemhandouts.html limiting reagents theoretical actual and percent yields example of a limiting reagent problem how many grams of NH<sub>3</sub> can be produced theoretically from the reaction of 5.0 g of N  
 what is the limiting reagent if 8.52 g are actually formed what is the

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 step one figure out how many moles of the limiting reagent you have  
 $\frac{21.3 \text{ grams O}_2}{32 \text{ grams/mole}} = 0.6656 \text{ moles O}_2$

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web percent yield is a measure of the efficiency of the experimental design  
 yield =  $\frac{\text{mass of product obtained}}{\text{calculated mass of the product expected}} \times 100$   
 in this reaction a yield of 80% is anticipated

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web a write a balanced equation for the reaction Al + O<sub>2</sub> → Al<sub>2</sub>O<sub>3</sub>  
 4 Al + 3 O<sub>2</sub> → 2 Al<sub>2</sub>O<sub>3</sub>  
 b determine the theoretical yield of Al<sub>2</sub>O<sub>3</sub>  
 $\frac{160.0 \text{ g Al}}{27 \text{ g/mol}} = 5.926 \text{ mol Al}$   
 $\frac{5.926 \text{ mol Al}}{4} \times 2 = 2.963 \text{ mol Al}_2\text{O}_3$   
 $2.963 \text{ mol Al}_2\text{O}_3 \times 101.96 \text{ g/mol} = 302.3 \text{ g Al}_2\text{O}_3$   
 c determine the percent yield  
 $\frac{260.0 \text{ g}}{302.3 \text{ g}} \times 100 = 86.01\%$   
 theoretical = 302.3 g

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