Period Date Tuesday, November 1 2016

Periodic Trends

ATOMIC RADIUS

What trend in atomic radius do you see as you go down a group/family on the periodic table?

What causes this trend?

3. What trend in atomic radius do you see as you go across a period/row on the periodic table?

4. What causes this trend?

Circle the atom in each pair that has the largest atomic radius.

a) Al B

b) S O

c) Br Cl

d) Na Al e) O F f) Mo Ca

Put the following elements in order from smallest to largest atomic radius and explain why: C. O. Sn. Sr.

ELECTRONEGATIVITY

Define electronegativity

8. How does the ionic radius of a nonmetal compare with its atomic radius?

9. What trend in electronegativity do you see as you go down a group/family on the periodic table?

10. What causes this trend?

11. What trend in electronegativity do you see as you go across a period/row on the periodic table?

12. What causes this trend?

Circle the atom in each pair that has the greater electronegativity.

a) Ca Ga

b) Li O c) Cl S d) Br As

e) Ba Sr

GENERAL QUESTIONS

14. Which group tends to form +1 ions?

15. Which group tends to form +2 ions?

16. Which group tends to form -1 ions?

17. Which group tends not to form ions or react?

Periodic Trends Worksheet Extension Questions Answers

Tobias Bleicker

Periodic Trends Worksheet Extension Questions Answers:

Merrill Chemistry Robert C. Smoot, Smoot, Richard G. Smith, Jack Price, 1998

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