

1. Circle the element with the largest atomic radius and put a square around the element with the smallest atomic radius:

Cu **K** Ni **Br**

- a. Explain why you made these choices: *All of the elements are in the same period. The trend in atomic radius as you go across a period is DECREASING. Therefore, the element on the far left (K) is the largest, and the element on the far right (Br) is the smallest.*

2. Circle the element with the highest ionization energy and put a square around the element with the lowest ionization energy:

Cu **K** Ni **Br**

- a. Explain why you made these choices: *All of the elements are in the same period. The trend in ionization energy as you go across a period is INCREASING. Therefore, the element on the far left (K) has the lowest ionization energy, and the element on the far right (Br) has the highest ionization energy.*

3. Circle the element with the highest electronegativity and put a square around the element with the lowest electronegativity:

Cu **K** Ni **Br**

- a. Explain why you made these choices: *All of the elements are in the same period. The trend in electronegativity as you go across a period is INCREASING. Therefore, the element on the far left (K) has the lowest electronegativity, and the element on the far right (Br) has the highest electronegativity.*

4. For each of the following groups: Circle the element with the largest atomic radius and put a square around the element with the smallest atomic radius:

- 5.
- O C **Be** **Ne** Same Period
 - Na Rb **Fr** **H** Same Group
 - Pb** **C** Sn Si Same Group
 - Au W S **Fr** **Ne** Zn Challenge

6. For each of the following groups: Circle the element with the highest ionization energy and put a square around the element with the lowest ionization energy:

- O C **Be** **Ne** Same Period
- Na Rb **Fr** **H** Same Group
- Pb** **C** Sn Si Same Group
- Au W S **Fr** **Ne** Zn Challenge

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David Baud



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Homenatge professor Josep M. Costa (eBook) 1a part. Trends in electrochemistry and corrosion at the beginning of the 21st century Pere-Lluís Cabot Julia, Enric Brillas Coso, 2004-04-13 Esta primera parte del libro Trends in Electrochemistry and Corrosion at the beginning of the 21st century dedicado al Prof Josep M Costa en ocasi n de su 70 aniversario re ne un total de 33 art culos y revisiones originales tanto cient ficas como tecnol gicas correspondientes al campo de la Electroqu mica Estos trabajos est n escritos en espa ol e ingl s por unos 110 investigadores de todo el mundo y son exponentes representativos del gran nivel alcanzado por la investigaci n internacional en diversas materias de gran inter s en la Electroqu mica de principios de este siglo XXI Los trabajos se han agrupado en 5 cap tulos generales relativos a los campos de Electroqu mica Molecular Electrodeposici n Electrodos Modificados Descontaminaci n Electroqu mica y Sensores y Electroan lisis This first part of the book Trends in Electrochemistry and Corrosion at the beginning of the 21st century dedicated to Professor Josep M Costa in occasion of his 70th birthday collects 33 original papers and reviews both scientific and technologic corresponding to the field of Electrochemistry These works are written in English and Spanish by about 110 researchers of all around the world and are representative exponents of the high level reached by the international research in several topics of great interest in Electrochemistry at the beginning of the 21st Century The works have been distributed into five general chapters related to the fields of Molecular Electrochemistry Electrodeposition Modified Electrodes Electrochemical Depollution and Sensors and Electroanalysis

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