

PERIODIC CLASSIFICATION OF ELEMENT

DOBERNIER TRIADS:-

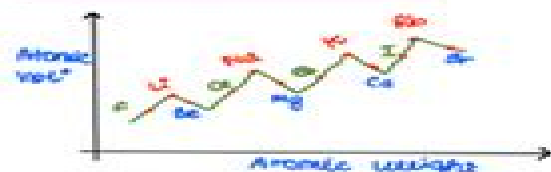
- Arranged elements in the group of 3.
- Atomic mass of middle element is average of atomic mass of other two elements.



Drawbacks:-

- as known elements were not grouped

Lothar Meyer's Curves:-



- Alkali metals → top position
- Alkaline earth metals → descending
- Halogen → ascending position.

MODERN PERIODIC TABLE:-

- long form of periodic table
- based on Moseley law: $\sqrt{\lambda} = a - b(Z - c)$
- Periodic properties are periodic function of their atomic number.

Periodicity - Repetition of similar properties of elements after regular interval of time.

Cause of periodicity - due to similar e^- configuration.

Transition elements: they have incompletely filled d-orbitals in their ground state.

Zn, Cd, Hg → are not transition elements because of completely filled d-orbitals. (pseudo transition elements).

Representative elements - Combined elements of s and p block are known as representative elements.

NEWLAND'S OCTAVE RULE:-

- When elements are arranged in increasing order of atomic mass, then every 8th element standing from first will show same property.



Drawbacks:-

1. Valid upto calcium only
2. No position of inert gases
3. Not worked on heavier elements.

MENDLEEV'S PERIODIC

- Periodic property of elements are periodic function of their atomic masses.

Merits:-

- Simplified study
- Some vacant space was left for the elements which were not known at that time.

Eka aluminium → Ga
Eka silicon → Ge

- Corrected atomic mass of element

Demerits:-

- (a) Position of hydrogen was not fixed
- (b) Could not explain periodicity.

MODERN PERIODIC TABLE:-

18 groups and 7 periods.

4-Blocks

s-Block = G_1 and $G_2 = ns^1-2$

p-Block = $G_{13} - G_{18} = ns^2 np^1-6$

d-Block = $G_3 - G_{10} = (n-1)d^1-10 ns^2$

f-Block = lanthanoid = $2 = 5f^1-7$

actinoid = $2 = 5f^1-10$

Diagonal Relation:

Li → Be → B → similar
Na → Mg → Al → change ratio
size

Nomenclature:

0 → nil	5 → pent	Suffix = turn
1 → un	6 → hex	100 unnilpentium
2 → bi	7 → sept	
3 → tri	8 → oct	
4 → quad	9 → enn	

EFFECTIVE NUCLEAR CHARGE

$Z_{eff} = Z - \sigma$

Trends - Period → Z_{eff} ↑

Group → Z_{eff} almost remains same. or Z_{eff} ↓↓

Shielding → $s > p > d > f$

→ more +ve charge, more is Z_{eff}

→ more -ve charge, less is Z_{eff} .

IONISATION ENERGY

→ Energy required to remove an electron from an isolated gaseous atom.

Mg → $M^{+}_{(g)} + e^-$ IE₁

$M^{+}_{(g)} \rightarrow M^{2+}_{(g)} + e^-$ IE₂

IE = IE₁ < IE₂ < IE₃ ...

Trends - Z_{eff} & IE

Period - Z_{eff} ↑ = IE ↑

Group - Z_{eff} ↓ = IE ↓

$G_{15} = G > Al > Ga > In$

$G_{14} = C > Si > Ge > Pb > Sn$

Period 2:-

$Li < Be < B < C < N < O < F < Ne$

ATOMIC RADIUS:-

- Distance b/w nucleus and electron of outermost shell.
- But it's impossible to find exact atomic radii.

Reason-1 = atoms too small

2 = atom doesn't have any boundary

Covallent Radius - Half of the internuclear distance b/w 2 atoms bonded by covalent bond.



Metallic Radius - Half of the internuclear distance b/w 2 atoms bonded by a metallic bond.



Van der Waals Radius - Half of the internuclear distance b/w 2 nearest molecules in non-bonded state.



Trends: Atomic Radius ∝ $1/Z_{eff}$
Period → Z_{eff} ↑ = Atomic Radius ↓
Group → Z_{eff} ↓ = Atomic Radius ↑

Ionic Radius + Radius < Neutral Radius

ELECTRON GAIN ENTHALPY:-

→ Energy required or released when an e^- is added to outermost orbit of an isolated gaseous atom. It is known as Electron gain enthalpy.



Note - In majority cases $\Delta_{eg}H_1$ is -ve (exothermic)

Exception = N } $\Delta_{eg}H_2$ = +ve
 noble gas } due to stable configuration

Note - Successive electron gain enthalpy is always +ve.

Trends: $\Delta_{eg}H \propto Z_{eff}$

Period - Z_{eff} ↑ → $\Delta_{eg}H$ ↑

Group - Z_{eff} ↓ → $\Delta_{eg}H$ ↓

$G_{15} = S > Se > Te > Po > O$
→ due to small size Reason!
→ e^-e^- repulsion.

Periodic Classification 3rd Chapter 11th Class

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 Liquids Unit VI Chemical Thermodynamics Unit VII Equilibrium Unit VIII Redox Reactions Unit IX Hydrogen Unit X s Block
 Elements Alkali and Alkaline earth metals Group 1 and Group 2 Elements Unit XI Some p Block Elements General
 Introduction to p Block Elements Unit XII Organic Chemistry Some Basic Principles and Techniques Unit XIII Hydrocarbons
 Classification of Hydrocarbons Unit XIV Environmental Chemistry Content 1 Some Basic Concepts of Chemistry 2 Structure
 of Atom 3 Classification of Elements and Periodicity in Properties 4 Chemical Bonding and Molecular Structure 5 States of
 Matter 6 Thermodynamics 7 Equilibrium 8 Redox Reactions 9 Hydrogen 10 s Block Elements 11 p Block Elements 12 Organic
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Table of Contents Periodic Classification 3rd Chapter 11th Class

1. Understanding the eBook Periodic Classification 3rd Chapter 11th Class
 - The Rise of Digital Reading Periodic Classification 3rd Chapter 11th Class
 - Advantages of eBooks Over Traditional Books
2. Identifying Periodic Classification 3rd Chapter 11th Class
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Periodic Classification 3rd Chapter 11th Class
 - User-Friendly Interface
4. Exploring eBook Recommendations from Periodic Classification 3rd Chapter 11th Class
 - Personalized Recommendations

- Periodic Classification 3rd Chapter 11th Class User Reviews and Ratings
- Periodic Classification 3rd Chapter 11th Class and Bestseller Lists
- 5. Accessing Periodic Classification 3rd Chapter 11th Class Free and Paid eBooks
 - Periodic Classification 3rd Chapter 11th Class Public Domain eBooks
 - Periodic Classification 3rd Chapter 11th Class eBook Subscription Services
 - Periodic Classification 3rd Chapter 11th Class Budget-Friendly Options
- 6. Navigating Periodic Classification 3rd Chapter 11th Class eBook Formats
 - ePub, PDF, MOBI, and More
 - Periodic Classification 3rd Chapter 11th Class Compatibility with Devices
 - Periodic Classification 3rd Chapter 11th Class Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Periodic Classification 3rd Chapter 11th Class
 - Highlighting and Note-Taking Periodic Classification 3rd Chapter 11th Class
 - Interactive Elements Periodic Classification 3rd Chapter 11th Class
- 8. Staying Engaged with Periodic Classification 3rd Chapter 11th Class
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Periodic Classification 3rd Chapter 11th Class
- 9. Balancing eBooks and Physical Books Periodic Classification 3rd Chapter 11th Class
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Periodic Classification 3rd Chapter 11th Class
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Periodic Classification 3rd Chapter 11th Class
 - Setting Reading Goals Periodic Classification 3rd Chapter 11th Class
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Periodic Classification 3rd Chapter 11th Class
 - Fact-Checking eBook Content of Periodic Classification 3rd Chapter 11th Class

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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