

Chapter 1: Limiting Reagents and Percent Yield (Pages 1-4)

Remember: Limiting reagent is the reactant that will completely react and leave another reactant leftover. **Remember:** Percent yield is the amount of product that is actually produced divided by the theoretical yield, multiplied by 100.

Remember: Limiting reagent is the reactant that will completely react and leave another reactant leftover. **Remember:** Percent yield is the amount of product that is actually produced divided by the theoretical yield, multiplied by 100.



The reaction is balanced. The limiting reagent is the reactant that will completely react and leave another reactant leftover.

2H ₂	20.0 g
O ₂	16.0 g
H ₂ O	36.0 g
H ₂	20.0 g

1. 20.0 g of H₂ and 16.0 g of O₂ are reacted. The limiting reagent is H₂. The theoretical yield of H₂O is 36.0 g.



2. The limiting reagent is H₂. The theoretical yield of H₂O is 36.0 g.

3. The limiting reagent is H₂. The theoretical yield of H₂O is 36.0 g.



4. The limiting reagent is H₂. The theoretical yield of H₂O is 36.0 g.



Percent Yield And Limiting Reactants Quiz

R Barnett



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