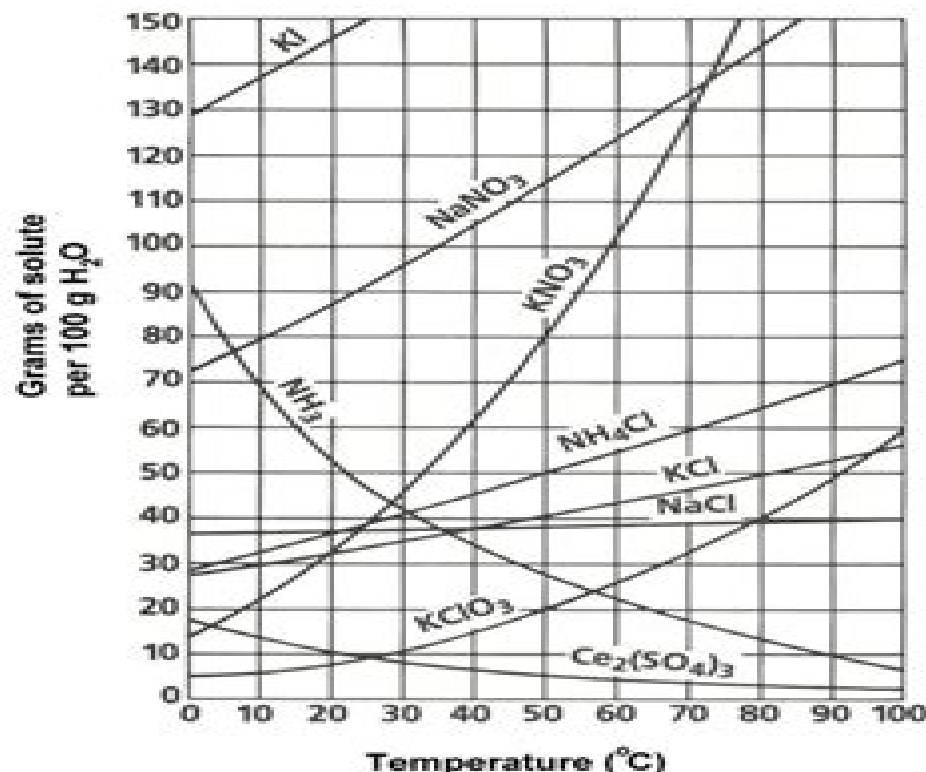


### Reading a Solubility Chart

- 1) The **curve** shows the # of grams of solute in a **saturated solution** containing 100 mL or 100 g of water at a certain temperature. (remember: 1 mL H<sub>2</sub>O = 1 g H<sub>2</sub>O)
- 2) Any amount of solute **below** the line indicates an **unsaturated solution** at a certain temperature
- 3) Any amount of solute **above** the line in which all of the solute has dissolved indicates a **supersaturated solution**.
- 4) If the amount of solute is above the line but has not all dissolved, the solution is saturated and the # grams of solute settled on the bottom of the container = total # g in solution – # g of a saturated solution at that temperature. (according to the curve)
- 5) Solutes whose curves move upward w/ increased temperature are typically solids b/c the solubility of solids increases w/ increased temperature.
- 6) Solutes whose curves move downward w/ increased temperature are typically gases b/c the solubility of gases decreases with increased temperature.



### Solubility Problems example (Many may be done by reading the solubility table)

1. At 10°C, 80 g of NaNO<sub>3</sub> will dissolve in 100 mL (a saturated solution)
2. To find the # grams needed to saturate a solution when the volume is NOT 100 mL use the following strategy to find answer: (e. g., 60 mL of H<sub>2</sub>O will be saturated with how many grams of NaNO<sub>3</sub> ?)

Start w/ known vol. x Solubility/100mL at set temp. = amount of Solute needed to saturate

Ex. 60 mL H<sub>2</sub>O x  $\frac{80 \text{ g NaNO}_3}{100 \text{ mL H}_2\text{O}}$  = 48 g NaNO<sub>3</sub> needed to saturate solution

or if the chart is in units of 100 g of H<sub>2</sub>O use the density of water conversion 1mL H<sub>2</sub>O= 1 g H<sub>2</sub>O

Ex. 60 mL H<sub>2</sub>O x  $\frac{1 \text{ g H}_2\text{O}}{1 \text{ mL H}_2\text{O}}$  x  $\frac{80 \text{ g NaNO}_3}{100 \text{ g H}_2\text{O}}$  = 48 g NaNO<sub>3</sub>

# Reading The Solubility Chart Problems Answer Key

**Marco Cascella**



## Reading The Solubility Chart Problems Answer Key:

Frank Schaffer's Chemistry for Everyday David Thurllo, 1999 Students learn about important subjects by relating them to events and things that occur in their everyday lives A wealth of interesting activities provide a detailed look into each subject Easy to use activities can be completed individually at school or at home though a few hands on experiments require group work and data sharing A great supplement to any existing curriculum Includes topics such as the scientific method applied to chemistry determining specific gravity balancing chemical equations and exploring the periodic table of elements

**Transactions** American Institute of Mining, Metallurgical, and Petroleum Engineers, 1938 Some vols 1920 1949 contain collections of papers according to subject *Transactions* Metallurgical Society of AIME., 1938 Papers and Discussions Presented Before the [Coal] Division American Institute of Mining, Metallurgical, and Petroleum Engineers, 1938

*Transactions of the American Institute of Mining and Metallurgical Engineers* , 1938 *British Books in Print* , 1970  
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## **Reading The Solubility Chart Problems Answer Key** Book Review: Unveiling the Power of Words

In a global driven by information and connectivity, the power of words has become more evident than ever. They have the ability to inspire, provoke, and ignite change. Such may be the essence of the book **Reading The Solubility Chart Problems Answer Key**, a literary masterpiece that delves deep into the significance of words and their affect our lives. Published by a renowned author, this captivating work takes readers on a transformative journey, unraveling the secrets and potential behind every word. In this review, we shall explore the book is key themes, examine its writing style, and analyze its overall impact on readers.

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