CHAPTER 6 STUDY GUIDE

The Periodic Table and Periodic Law

Section 6.1 Development of the Modern Periodic Table

In your textbook, reads about the history of the periodic table's development.

Use each of the terms below just once to complete the passage.

| octaves | atomic mass | atomia-number | The state of the s |
|-----------|---------------------|------------------|--|
| colomerus | <u>properties -</u> | Henry Messley | |
| protons | periodic law | Desira Mandalosv | accented |

| 1 2 | 3 | 4 | 5 | 6 | 7 | |
|-----|----|----|------|----|----|-----|
| н | Li | G | Bo | C | N | 0 |
| 8 | 9 | 10 | TI | 12 | 13 | 1.5 |
| E | Na | Mg | : AJ | Si | P | S |

The first periodic table is mostly credited to (5) Mendelety In his table, the elements were arranged according to increasing (6) Office Mass. One important result of this table was that the existence and properties of undiscovered

(7) Elements could be predicted.

The element in the modern periodic table are arranged according to increasing

(8) Offmic number, as a result of the work of (9) Moseloy. This arrangement is based on number of (10) Profos in the nucleus of an atom of the element. The modern form of the periodic table results in the (11) Periodic last. which states that when elements are arranged according to increasing atomic number, there is a periodic repetition of their chemical and physical (12) Properties.

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