

Chemistry - Unit 1 Chapters 1 & 2**Element Quiz:****Problem Set Due:****Vocab quiz:****Test Date:****Vocabulary:** Posted online – there will be a matching vocab quiz

- | | | | | | |
|------------------------|--------------------------|-------------------|-------------------|---------------|-------------|
| 1 matter | 2 temperature | 3 physical change | 4 chemical change | 5 precipitate | 6 chemistry |
| 7 compounds | 8 elements | 9 solution | 10 quantitative | 11 weight | 12 mass |
| 13 length | 14 SI | 15 accuracy | 16 precision | 17 density | 18 volume |
| 19 significant figures | 20 heterogeneous mixture | 21 endothermic | 22 exothermic | | |

Formulas/constants to be memorized:metric conversions given elements from handout
density(packet)

percent error (given)

Conversions that do NOT have to be memorized:

1 lb = 453.6 g 1 mile = 1.609 km 1 L = 1.0567 qts 1 oz = 28.35 g 4 qts = 1 gal
 1 m = 39.37 in = 3.2808 ft = 1.0936 yd (all equal! Use any 2 in a conversion) 1 fluid oz = 29.6 mL

OBJECTIVES:

- Be able to apply the concept of significant digits when recording lab data
- Memorize the metric system; be able to do both pure metric conversions and English-metric conversion. (You will be given English-metric conversions).
- Be able to calculate percent error from lab data.
- Use factor label (dimensional analysis) to solve math problems and conversions.
- Be able to do density problems.
- Be able to explain the difference between accuracy and precision and weight and mass.
- Be able to use the above vocabulary properly.
- Be able to classify properties and changes as physical or chemical.
- Understand and apply basic laboratory safety rules.
- Be able to visually identify and correctly use lab equipment and safety equipment.
- Know the fire escape route for the classroom and other safety procedures.
- Be able to collect, assimilate, and draw conclusions from scientific data.

Unit 1BQs We will work on this together in class! If you are absent you must still complete the BQs!! LABEL Q with numbers circle your final answers! Include units on your answers. 2. Show your work on all problems. Credit will not be given without work.

- Write the symbol or name for the following elements: K, Na, carbon, sulfur, nitrogen, lead and F
- In lab you find the mass of an object to be 5.0 g the correct mass is 5.2 g - calculate your percent error.
- Id the following as physical or chemical changes:
A. Wood burning in a bonfire B. inflating a tire with air C. ice cream cone melting D. sugar dissolving in water E. water absorbed by a paper towel F. your car using gasoline
- ID the following as heterogeneous mixtures, elements, solutions or compounds:
A. Kool-Aid B. ice C. pure gold coin D. vegetable soup E. diamond F. salt
- How many quarters in \$10?
- Convert: A. 55 meters to km B. 25 mm to cm C. 252 g to lb
- List two common sense lab safety rules.



- Record the length of the paper for A and B. which ruler is more accurate? Why? (the unit for both rulers is cm).
- What is the density of 20.0 g of a solution that has a volume of 38.2 ml?
- if you are traveling 25 ft/sec A. What is your speed in m/min ? B. How many km can you travel in half an hour?
- When a sample of a metal that has a mass of 15.8 g is placed into a graduated cylinder containing water, the volume increases from 16.0 ml to 19.5 ml.
A. What is the volume of the substance? B. What is the density of the metal?
C. Using your reference packet identify the metal.

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