

**Chapter 12: Atoms and the Periodic Table**  
**Subatomic Particles**

Three fundamental particles of the atom are the \_\_\_\_\_, \_\_\_\_\_, and \_\_\_\_\_. At the center of each atom lies the atomic \_\_\_\_\_, which consists of \_\_\_\_\_ and \_\_\_\_\_. The atomic number refers to the number of \_\_\_\_\_ in the nucleus. All atoms of the same element have the same number of \_\_\_\_\_, hence, the same atomic number.

Isotopes are atoms that have the same number of \_\_\_\_\_ but a different number of \_\_\_\_\_. An isotope is identified by its atomic mass number, which is the total number of \_\_\_\_\_ and \_\_\_\_\_ in the nucleus. A carbon isotope that has 6 \_\_\_\_\_ and 6 \_\_\_\_\_ is identified as carbon-12, where 12 is the atomic mass number. A carbon isotope having 6 \_\_\_\_\_ and 8 \_\_\_\_\_, on the other hand, is carbon-14.

1. Complete the following table:

Isotope	Number of...		
	Electrons	Protons	Neutrons
Hydrogen-1	1		
Chlorine-36		17	
Nitrogen-14			7
Potassium-40	19		
Arsenic-75		33	
Gold-197			118

2. Which results in a more valuable product — adding or subtracting protons from gold nuclei?

3. Which has more mass, a helium atom or a neon atom?

4. Which has a greater number of atoms, a gram of helium or a gram of neon?



# Physical Science Chapter 12 Answers

**Jin-Ying Zhang**



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*Glencoe Physical Science*, 1999

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