

## Chapter 1 Science Skills

**WordWise**

Answer the questions by writing the correct vocabulary term in the blanks. Use the circled letter in each term to find the hidden vocabulary word. Then write a definition for the hidden word.

**Clues**

The study of matter, energy, and their interactions

The closeness of a measurement to the actual value of what is being measured

A gauge of how exact a measurement is

The ratio of a vertical change to the corresponding horizontal change in a line

An instrument used to measure temperature

The use of knowledge to solve practical problems

A representation of an object or event

A system of knowledge and the methods used to find that knowledge

A statement that summarizes a pattern found in nature

Information that you obtain through your senses

**Hidden word:** \_\_\_\_\_

**Definition:** \_\_\_\_\_

**Vocabulary Terms**

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# Physical Science Ch 12 Wordwise Answers

**Patrick Vollmar**



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web 1 answer a explanation the value of  $\Delta h$  is negative therefore the reaction is exothermic or heat releasing 2 answer a 1 explanation the entire reaction was multiplied by 2 therefore  $\Delta h$  must also be halved 3 answer c explanation option c is incorrect because 566 kj mol is the  $\Delta h$  per mole of  $\text{O}_2$  and not for 2 moles

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web thermochemical equations equations that represent both mass and enthalpy changes  $\text{H}_2\text{O}(\text{s}) \rightarrow \text{H}_2\text{O}(\text{l})$   $\Delta h = 6.01 \text{ kJ/mol}$  this is an endothermic process it requires 6.01 kJ to melt one mole of ice  $\text{H}_2\text{O}(\text{s})$  the enthalpy value will change if the number of moles varies from the 1 1 reaction stoichiometry this is an exothermic process

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web in this set of practice questions we will summarize the main concepts of thermochemistry such as the relationship between internal energy work and heat exothermic and endothermic process heat capacity constant pressure calorimetry constant volume calorimetry the enthalpy the standard enthalpies of formation and their use in

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web we will revisit this today and go on to see that if we use a special kind of thermochemical reaction called the standard

enthalpy of formation we can calculate enthalpies of reactions without having to manipulate a series of individual thermochemical equations for each step

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web we can express the first law in many ways one of the more useful expressions is that the change in internal energy  $e$  of a system in any process is equal to the heat  $q$  added to the system plus the work  $w$  done on the system by its surroundings  $e$

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