

QUESTION 2

- 2.1 The graph of molecular size versus the boiling point is given below.
The letters **A**, **B** and **C** represent the compounds CH_4 , C_2H_6 and C_3H_8 respectively.



- 2.1.1 Define the term *boiling point*. (2)
- 2.1.2. Describe the trend in the boiling points of the compounds as shown by the graph. (2)
- 2.1.3. Explain the answer to QUESTION 2.1.2 by referring to MOLECULAR SIZE, TYPE and STRENGTH of INTERMOLECULAR FORCES. (3)
- 2.1.4 Which ONE of the compounds (**A**, **B** or **C**) has the HIGHEST vapour pressure? Explain the answer by referring to the data on the graph. (2)

- 2.2 Consider the two molecules in the table below.

| NAME OF SUBSTANCE | FORMULA | MOLACULAR MASS (g) | BOILING POINT ($^{\circ}\text{C}$) |
|-------------------|---------------|--------------------|--------------------------------------|
| Ammonia | NH_3 | 17 | - 33 |
| Phosphine | PH_3 | x | - 87,4 |

- 2.2.1 Determine the molecular mass (x) of phosphine (1)
- 2.2.2 Explain the difference in the boiling points by referring to the TYPE and STRENGTH of INTERMOLECULAR FORCES. (3)

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Table of Contents Physical Science Grade 1march Question Papers

1. Understanding the eBook Physical Science Grade 1march Question Papers
 - The Rise of Digital Reading Physical Science Grade 1march Question Papers
 - Advantages of eBooks Over Traditional Books
2. Identifying Physical Science Grade 1march Question Papers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Physical Science Grade 1march Question Papers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Physical Science Grade 1march Question Papers
 - Personalized Recommendations
 - Physical Science Grade 1march Question Papers User Reviews and Ratings
 - Physical Science Grade 1march Question Papers and Bestseller Lists

5. Accessing Physical Science Grade 1march Question Papers Free and Paid eBooks
 - Physical Science Grade 1march Question Papers Public Domain eBooks
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 - Physical Science Grade 1march Question Papers Budget-Friendly Options
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 - Physical Science Grade 1march Question Papers Enhanced eBook Features
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 - Adjustable Fonts and Text Sizes of Physical Science Grade 1march Question Papers
 - Highlighting and Note-Taking Physical Science Grade 1march Question Papers
 - Interactive Elements Physical Science Grade 1march Question Papers
8. Staying Engaged with Physical Science Grade 1march Question Papers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Physical Science Grade 1march Question Papers
9. Balancing eBooks and Physical Books Physical Science Grade 1march Question Papers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Physical Science Grade 1march Question Papers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Physical Science Grade 1march Question Papers
 - Setting Reading Goals Physical Science Grade 1march Question Papers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Physical Science Grade 1march Question Papers
 - Fact-Checking eBook Content of Physical Science Grade 1march Question Papers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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