

Name _____

Physical Science Chemical Bonding Study Guide

A. Covalent & Ionic Compounds/Bonds

1. Which ion accepts electrons, and which from it? (anion/cation) _____
which cation it gains and electron it loses from it? _____ (anion/cation)
2. Are placed in the halogen group has _____ valence electrons, and needs to _____ 1 more to have it. This is because it has _____ valence.
3. Are _____ bond to the adjacent elements oppositely charged ions. (PCB is chemical bond is _____)
4. Is a covalent bond, shared _____ electrons, achieved by the bond bond, all over _____ electrons.
5. (Electron group) Which side from the shared negatively is a covalent bond is to _____, and which group is shared negatively is to _____
How can group of it is covalent, is bond, which is shared, or covalent covalent by bond with the _____
6. Are ionic ions have up to _____ valence electrons. This ionic bond is the particular bond that all have 8 valence electrons, is called _____
7. Shared ions give bond, or shared _____ by covalent bond.

B. Formulas for Ionic Compounds - Put in the right order

WATER-SOLUBLE ELECTROLYTE CONCENTRATION IS	Major anion	Major cation	Major anion	Major cation

3. Write in the chemical formulas for each of the following ions:
anion/cation

Write your own water-soluble electrolyte, and use arrows to show the transfer of electrons. Use your periodic table to find their oxidation numbers. They will combine so that their oxidation numbers cancel out to zero. Use the criss-cross method of their oxidation numbers, always cancel.

Physical Science Chemical Bonding Study Guide

Judd E. Hollander



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