

### Chapter 6: Chemical Bonds Review Answer Key

- What is the most stable group on the Periodic Table? Why is this group stable?  
Noble Gases (Group 18) because they have 8 valence electrons which means their outer energy level is full.
- Name the most reactive group(s) on the Periodic Table? Why are they reactive?  
Alkali Metals (Group 1) & Halogens (Group 17): Alkali Metals want to lose one valence electron in order to become stable and Halogens want to gain one valence electron to be stable.

- In the boxes below draw the Electron dot diagrams for:

| Sodium            | Calcium                      | Boron                             | Chlorine                                                                                                              | Krypton                                                                                                               |
|-------------------|------------------------------|-----------------------------------|-----------------------------------------------------------------------------------------------------------------------|-----------------------------------------------------------------------------------------------------------------------|
| $\text{Na} \cdot$ | $\text{Ca} \cdot$<br>$\cdot$ | $\cdot \text{B} \cdot$<br>$\cdot$ | $\begin{array}{c} \cdot \cdot \\ \cdot \cdot \\ \cdot \cdot \\ \cdot \cdot \\ \cdot \cdot \\ \cdot \cdot \end{array}$ | $\begin{array}{c} \cdot \cdot \\ \cdot \cdot \\ \cdot \cdot \\ \cdot \cdot \\ \cdot \cdot \\ \cdot \cdot \end{array}$ |

- Name the three (3) types of chemical bonds:  
Ionic bonds      Covalent bonds      Polyatomic
- Define ionic bond:  
The force that holds cations and anions together.
- Define covalent bond:  
A chemical bond in which two atoms share a pair of valence electrons.
- Define polyatomic ion:  
A covalently bonded group of atoms that has a positive or negative charge and acts as a unit.
- In an ionic bond, electrons are transferred.
- Who are the participants in an ionic bond?      Metals and nonmetals
- An ion with a positive charge is a cation; an ion with a negative charge is an anion.
- In a covalent bond, electrons are shared.
- Who are the participants in a covalent bond? Two nonmetals
- Identify whether the following compounds are covalent (two nonmetals), ionic (a nonmetal and a metal), or polyatomic (both).
  - $\text{CaCl}_2$       IONIC (transferred)
  - $\text{H}_2\text{O}$       COVALENT (shared)
  - $\text{BaSO}_4$       POLYATOMIC (both)
  - $\text{CH}_4$       COVALENT (shared)
  - $\text{LiBr}$       IONIC (transferred)
  - $\text{NH}_4\text{Cl}$       POLYATOMIC (both)
  - $\text{KI}$       IONIC (transferred)

# Physical Science Chapter 6 Chemical Bonds Wordwise Answers

**D Kirk**



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