

All quantities on the right are given to us except n , which can be computed from the given mass of Ar.

$$n = \frac{25 \text{ g}}{39.95 \text{ g mol}^{-1}} = 0.626 \text{ mol}$$

$$\text{so } p = \frac{(0.626 \text{ mol}) \times (8.31 \times 10^{-2} \text{ dm}^3 \text{ bar K}^{-1} \text{ mol}^{-1}) \times (30 + 273 \text{ K})}{1.5 \text{ dm}^3} = \boxed{10.3 \text{ bar}}$$

not 2.0 bar.

(b) The van der Waals equation [1.21a] is

$$p = \frac{RT}{V_m - b} - \frac{a}{V_m^2}$$

$$\begin{aligned} \text{so } p &= \frac{(8.31 \times 10^{-2} \text{ dm}^3 \text{ bar K}^{-1} \text{ mol}^{-1}) \times (30 + 273) \text{ K}}{(1.5 \text{ dm}^3 / 0.626 \text{ mol}) - 3.20 \times 10^{-2} \text{ dm}^3 \text{ mol}^{-1}} \\ &\quad - \frac{(1.337 \text{ dm}^6 \text{ atm mol}^{-2}) \times (1.013 \text{ bar atm}^{-1})}{(1.5 \text{ dm}^3 / 0.626 \text{ mol})^2} = \boxed{10.4 \text{ bar}} \end{aligned}$$

E1.2(b) (a) Boyle's law [1.5] applies.

$$pV = \text{constant} \quad \text{so} \quad p_f V_f = p_i V_i$$

This equation can be solved for either initial or final pressure, hence

$$p_i = \frac{p_f V_f}{V_i} = \frac{(1.97 \text{ bar}) \times (2.14 \text{ dm}^3)}{(2.14 + 1.80) \text{ dm}^3} = \boxed{1.07 \text{ bar}}$$

(b) The original pressure in Torr is

$$p_i = (1.07 \text{ bar}) \times \left(\frac{1 \text{ atm}}{1.013 \text{ bar}} \right) \times \left(\frac{760 \text{ Torr}}{1 \text{ atm}} \right) = \boxed{803 \text{ Torr}}$$

E1.3(b) The relationship between pressure and temperature at constant volume can be derived from the perfect gas law, $pV = nRT$ [1.8]

$$\text{so } p \propto T \quad \text{and} \quad \frac{p_i}{T_i} = \frac{p_f}{T_f}$$

The final pressure, then, ought to be

$$p_f = \frac{p_i T_f}{T_i} = \frac{(125 \text{ kPa}) \times (11 + 273) \text{ K}}{(23 + 273) \text{ K}} = \boxed{120 \text{ kPa}}$$

E1.4(b) According to the perfect gas law [1.8], one can compute the amount of gas from pressure, temperature, and volume.

$$pV = nRT$$

$$\text{so } n = \frac{pV}{RT} = \frac{(1.00 \text{ atm}) \times (1.013 \times 10^5 \text{ Pa atm}^{-1}) \times (4.00 \times 10^3 \text{ m}^3)}{(8.3145 \text{ J K}^{-1} \text{ mol}^{-1}) \times (20 + 273) \text{ K}} = 1.66 \times 10^5 \text{ mol}$$

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CO Houle



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 - The Rise of Digital Reading Physical Chemistry 9th Solution Manual
 - Advantages of eBooks Over Traditional Books
2. Identifying Physical Chemistry 9th Solution Manual
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Physical Chemistry 9th Solution Manual
 - User-Friendly Interface
4. Exploring eBook Recommendations from Physical Chemistry 9th Solution Manual
 - Personalized Recommendations
 - Physical Chemistry 9th Solution Manual User Reviews and Ratings
 - Physical Chemistry 9th Solution Manual and Bestseller Lists
5. Accessing Physical Chemistry 9th Solution Manual Free and Paid eBooks
 - Physical Chemistry 9th Solution Manual Public Domain eBooks
 - Physical Chemistry 9th Solution Manual eBook Subscription Services
 - Physical Chemistry 9th Solution Manual Budget-Friendly Options
6. Navigating Physical Chemistry 9th Solution Manual eBook Formats

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- Physical Chemistry 9th Solution Manual Compatibility with Devices
- Physical Chemistry 9th Solution Manual Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Physical Chemistry 9th Solution Manual
 - Highlighting and Note-Taking Physical Chemistry 9th Solution Manual
 - Interactive Elements Physical Chemistry 9th Solution Manual
- 8. Staying Engaged with Physical Chemistry 9th Solution Manual
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Physical Chemistry 9th Solution Manual
- 9. Balancing eBooks and Physical Books Physical Chemistry 9th Solution Manual
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Physical Chemistry 9th Solution Manual
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Physical Chemistry 9th Solution Manual
 - Setting Reading Goals Physical Chemistry 9th Solution Manual
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Physical Chemistry 9th Solution Manual
 - Fact-Checking eBook Content of Physical Chemistry 9th Solution Manual
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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web the flymo xl500 petrol power air cushion mower is a durable practical and powerful mower designed for both domestic and professional usage with all the advantages of air cushion mowing afforded without sacrificing the benefits of petrol power mowing

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web list of flymo products flymo has sold various lawn tractors and push mowers in addition to other lawn and garden equipment other products include petrol lawnmowers hovering and wheeled wheeled electric lawnmowers hand mowers hedge trimmers edgers garden vacuum cleaners robotic lawnmowers references