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web an solution initially contains a catalytic amount of an enzyme with $k_m = 1.5 \text{ mmol/L}$ of substrate and no product after 45 seconds the solution contains $25 \mu\text{M}$ of product find v_{max} and the concentration of product after 20 minutes hint $s = k_m$
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web a reaction and its experimental rate law are represented below $2 \text{ICl} + \text{H}_2 \rightarrow 2 \text{HI} + 2 \text{HCl}$ rate $k [\text{ICl}]^2 [\text{H}_2]$ a chemist proposes a mechanism for the reaction that is consistent with the rate law the mechanism has two elementary steps and the first step is slow compared to the second

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 $\text{O}_3 \rightarrow \text{O}_2 + \text{O}$ fast
 $\text{O} + \text{O}_3 \rightarrow \text{O}_2 + \text{O}_2$ slow
 a write the overall balanced chemical equation
 $2 \text{ O}_3 \rightarrow 3 \text{ O}_2$ b identify any intermediates within the mechanism
 c what is the order with respect

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consider the following mechanism
 $\text{O}_3 \rightarrow \text{O}_2 + \text{O}$ fast
 $\text{O}_3 + \text{O} \rightarrow 2\text{O}_2$ slow
a) write the overall balanced chemical equation
b) identify any intermediates within the mechanism

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 $\text{Mg} + 2\text{HCl} \rightarrow \text{MgCl}_2 + \text{H}_2$
g the reaction between Mg and HCl is represented by the equation above in a kinetics experiment a 0.080 g sample of Mg ribbon is added to an erlenmeyer flask containing 50 ml of 3.0 M HCl at 25 °C

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