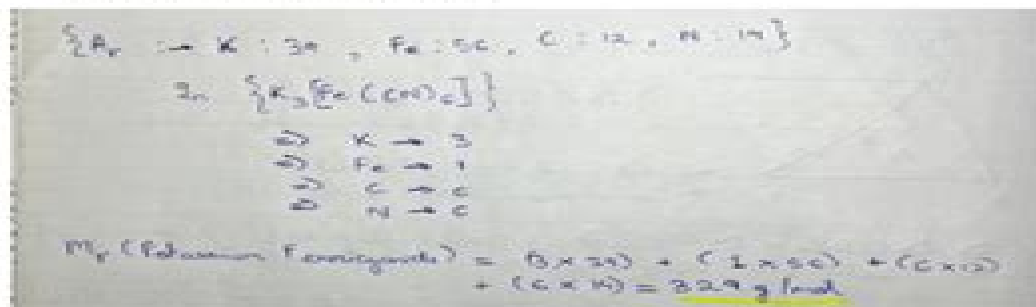


IGCSE CHEMISTRY (0620) REVISION NOTES

Chapter → Stoichiometry

1. In a compound, the elements that make up the compound are always combined in a specific ratio by mass. Say for instance, in the compound MgO , 0.24g of Mg will always combine with 0.16g of Oxygen. This can be extrapolated to any value. For example, 0.48g of Magnesium would always combine with 0.32g of Oxygen and so on.
2. A link between the chemical formula of a compound and their mass can be established. This is done using the concept of moles.
3. **The moles concept →**
 - a. Moles are a standardized counting unit
 - b. 1 mole of a substance contains approximately 6.02×10^{23} particles (atoms, molecules or formula units); this number is called the Avogadro's constant and has been given the symbol of " L ".
 - c. The mass of 1 mole of a substance is called the relative molecular mass of the substance (also known as molar mass); it is denoted as M_r . This value can be calculated by aggregating all the individual A_r value of each element in the compound
 - d. Process of calculating molar mass →
consider the process of calculating the M_r of Potassium Ferricyanide [$\text{K}_3\text{Fe}(\text{CN})_6$]



4. **Using the moles concept, we derive the following equations →**
 - a. Molar mass = mass of 1 mole = mass / moles
 - b. Number of particles = moles * Avogadro's constant

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E Durkheim



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