#### Six Types of Chemical Reaction Worksheet Answers

Balance the following reactions and indicate which of the six types of chemical reaction are being represented:

a) 2 NaBr + 1 Ca(OH)₂ → 1 CaBr₂ + 2 NaOH

Type of reaction: double displacement

b) 2 NH<sub>3</sub>+ 1 H<sub>2</sub>SO<sub>4</sub> → 1 (NH<sub>4</sub>)<sub>2</sub>SO<sub>4</sub>

Type of reaction: synthesis

c) 4 C<sub>8</sub>H<sub>9</sub>O + 27 O<sub>2</sub> → 20 CO<sub>2</sub> + 18 H<sub>2</sub>O

Type of reaction: combustion

d) 3 Pb + 2 H<sub>3</sub>PO<sub>4</sub> → 3 H<sub>2</sub> + 1 Pb<sub>3</sub>(PO<sub>4</sub>)<sub>2</sub>

Type of reaction: single displacement

e) 1 Li<sub>2</sub>N + 3 NH<sub>4</sub>NO<sub>3</sub> → 3 LiNO<sub>3</sub> + 1 (NH<sub>4</sub>)<sub>3</sub>N

Type of reaction: double displacement

f) 3 HBr + 1 Al(OH)<sub>3</sub> → 3 H<sub>2</sub>O + 1 AlBr<sub>3</sub>

Type of reaction: acid-base (double displacement)

g) Na<sub>3</sub>PO<sub>4</sub> + 3 KOH → 3 NaOH + K<sub>3</sub>PO<sub>4</sub>

g) Najrou + 3 Non -> 3 Naon + Njrou

h) MgCl₂ + Li₂CO₃ → MgCO₃ + 2 LiCl

i) C<sub>8</sub>H<sub>12</sub> + 9 O<sub>2</sub> → 6 CO<sub>2</sub> + 6 H<sub>2</sub>O

i) Pb + FeSO<sub>4</sub> → PbSO<sub>4</sub> + Fe

k) CaCO<sub>3</sub> → CaO + CO<sub>2</sub>

P<sub>4</sub> + 3 O<sub>2</sub> → 2 P<sub>2</sub>O<sub>3</sub>

m) 2 RbNO<sub>3</sub> + BeF<sub>2</sub> → Be(NO<sub>3</sub>)<sub>2</sub> + 2 RbF

n) 2 AgNO<sub>3</sub> + Cu → Cu(NO<sub>3</sub>)<sub>2</sub> + 2 Ag

O) C<sub>3</sub>H<sub>8</sub>O + 4 O<sub>2</sub> → 3 CO<sub>2</sub> + 3 H<sub>2</sub>O

p) 2 CsHs + Fe → Fe(CsHs)<sub>2</sub>

q) SeCl<sub>0</sub> + O<sub>2</sub> → SeO<sub>2</sub> + 3Cl<sub>2</sub>

r) 2 Mgl<sub>2</sub> + Mn(SO<sub>3</sub>)<sub>2</sub> → 2 MgSO<sub>3</sub> + Mnl<sub>4</sub>

s) O<sub>2</sub> → O + O<sub>2</sub>

t) 2 NO<sub>2</sub> → 2 O<sub>2</sub> + N<sub>2</sub>

Reaction Type \_\_double displacement\_\_\_\_

Reaction Type double displacement

Reaction Type \_\_\_\_complete combustion\_\_\_\_

Reaction Type \_\_\_\_single displacement\_\_\_\_

Reaction Type \_\_\_\_decomposition\_\_\_\_\_

Reaction Type \_\_\_synthesis\_\_\_\_\_

Reaction Type \_\_\_\_double displacement\_\_\_\_

Reaction Type \_\_\_single displacement\_\_\_\_

Reaction Type \_\_\_\_complete combustion\_\_\_\_

Reaction Type \_\_\_synthesis\_\_\_\_\_

Reaction Type \_\_\_single displacement\_\_\_\_

Reaction Type \_\_\_\_double displacement\_\_\_\_

Reaction Type \_\_\_\_decomposition\_\_\_\_\_

Reaction Type \_\_\_\_\_decomposition\_\_\_\_\_

# **Reactions Classifications Answer Key**

**AN Whitehead** 

**Reactions Classifications Answer Key:** 

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## **Table of Contents Reactions Classifications Answer Key**

- 1. Understanding the eBook Reactions Classifications Answer Key
  - The Rise of Digital Reading Reactions Classifications Answer Key
  - Advantages of eBooks Over Traditional Books
- 2. Identifying Reactions Classifications Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Reactions Classifications Answer Key
  - User-Friendly Interface
- 4. Exploring eBook Recommendations from Reactions Classifications Answer Key
  - Personalized Recommendations
  - Reactions Classifications Answer Key User Reviews and Ratings
  - Reactions Classifications Answer Key and Bestseller Lists
- 5. Accessing Reactions Classifications Answer Key Free and Paid eBooks

- Reactions Classifications Answer Key Public Domain eBooks
- Reactions Classifications Answer Key eBook Subscription Services
- Reactions Classifications Answer Key Budget-Friendly Options
- 6. Navigating Reactions Classifications Answer Key eBook Formats
  - o ePub, PDF, MOBI, and More
  - Reactions Classifications Answer Key Compatibility with Devices
  - Reactions Classifications Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Reactions Classifications Answer Key
  - Highlighting and Note-Taking Reactions Classifications Answer Key
  - Interactive Elements Reactions Classifications Answer Key
- 8. Staying Engaged with Reactions Classifications Answer Key
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Reactions Classifications Answer Key
- 9. Balancing eBooks and Physical Books Reactions Classifications Answer Key
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Reactions Classifications Answer Key
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Reactions Classifications Answer Key
  - Setting Reading Goals Reactions Classifications Answer Key
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Reactions Classifications Answer Key
  - Fact-Checking eBook Content of Reactions Classifications Answer Key
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development

- Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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