

CHAPTER 10: Solutions

Reaction Energy

SECTION 1

Write a balanced chemical equation for the following reactions in the space provided.

1. For elements in their standard state, the value of ΔH_f° is _____.
2. The standard heat of decomposition is a measure for signification by the following thermodynamic equation:



- a. Describe the reaction between the two in terms of energy and enthalpy.
- b. What is the standard heat of decomposition for the reaction?

Write a balanced chemical equation for the reaction between the two in the space provided.

3. The standard heat of formation for CO_2 is -393.5 kJ/mol . Write a balanced chemical equation for the reaction between the two in the space provided.

4. The standard heat of formation for H_2O is -285.8 kJ/mol . Write a balanced chemical equation for the reaction between the two in the space provided.

5. The standard heat of formation for H_2O is -285.8 kJ/mol . Write a balanced chemical equation for the reaction between the two in the space provided.

- a. What is the value of ΔH_f° for this reaction?

Reaction Energy Section 1 Answer Key

Ian Pickup



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