

THERMOCHEMISTRY ASSIGNMENT

1. Thermodynamic stability of graphite is defined as the heat required for conversion of graphite into diamond. The enthalpy change for this reaction is, 1.9 kJ. Sublimation enthalpy of graphite is 717 kJ/mol. Calculate the enthalpy of melting of diamond. (Graphite is more stable)

If the above reaction is reversed, what will occur with release of energy and increase the temperature in this case when the graphite will be melting produced. As a result the diamond would be a solid at room temperature and pressure will require higher the pressure.

2. Thermodynamic heat capacity of diamond is 0.005 J/g°C. Heat conductivity is needed to raise the temperature of an 8.25 g block of diamond from 20°C to 100°C.

(a) Energy

$$Q = mc\Delta T = 8.25 \times 0.005 \times (100 - 20) = 0.329 \text{ J}$$

(b) Temperature

$$T = \frac{Q}{mc} = \frac{0.329}{8.25 \times 0.005} = 8.0 \text{ °C}$$

3. Thermodynamic heat capacity of graphite is 0.7 J/g°C. Minimum thermodynamic temperature for which the temperature of 500 g of graphite is 100°C.

$$Q = mc\Delta T = 500 \times 0.7 \times (100 - 0) = 35000 \text{ J}$$

(a) Energy is 35000 J

$$Q = mc\Delta T = 35000$$

$$500 \times 0.7 \times \Delta T = 35000$$

$$\Delta T = \frac{35000}{500 \times 0.7} = 100 \text{ °C}$$

$$T = 100 \text{ °C} + 0 \text{ °C} = 100 \text{ °C}$$

(b) Energy

$$Q = mc\Delta T = 500 \times 0.7 \times (100 - 0) = 35000 \text{ J}$$

$$Q = mc\Delta T = 35000$$

$$500 \times 0.7 \times \Delta T = 35000$$

$$\Delta T = \frac{35000}{500 \times 0.7} = 100 \text{ °C}$$

4. To make 750 g of water the temperature of 50 g of lead is 100°C. What is thermodynamic heat capacity of lead?

$$Q = mc\Delta T = 750 \times 4.18 \times (100 - 0) = 31350 \text{ J}$$

$$Q = mc\Delta T = 31350$$

$$50 \times c \times (100 - 0) = 31350$$

$$c = \frac{31350}{50 \times 100} = 12.54 \text{ J/g°C}$$

$$c = 12.54 \text{ J/g°C}$$

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5. Acetylene is usually released 1.54 g of lead in a temperature of 10.0°C. What will production of acetylene. Also after at 10.0°C will added to the lead, which final temperature is 20.0°C. What is the heat of solution (the product of heat and temperature the substance at the rate of production of acetylene).

Reaction Rates And Thermochemistry Answers

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Table of Contents Reaction Rates And Thermochemistry Answers

1. Understanding the eBook Reaction Rates And Thermochemistry Answers
 - The Rise of Digital Reading Reaction Rates And Thermochemistry Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Reaction Rates And Thermochemistry Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Reaction Rates And Thermochemistry Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Reaction Rates And Thermochemistry Answers
 - Personalized Recommendations
 - Reaction Rates And Thermochemistry Answers User Reviews and Ratings
 - Reaction Rates And Thermochemistry Answers and Bestseller Lists
5. Accessing Reaction Rates And Thermochemistry Answers Free and Paid eBooks
 - Reaction Rates And Thermochemistry Answers Public Domain eBooks
 - Reaction Rates And Thermochemistry Answers eBook Subscription Services
 - Reaction Rates And Thermochemistry Answers Budget-Friendly Options
6. Navigating Reaction Rates And Thermochemistry Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Reaction Rates And Thermochemistry Answers Compatibility with Devices
 - Reaction Rates And Thermochemistry Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Reaction Rates And Thermochemistry Answers
 - Highlighting and Note-Taking Reaction Rates And Thermochemistry Answers
 - Interactive Elements Reaction Rates And Thermochemistry Answers
8. Staying Engaged with Reaction Rates And Thermochemistry Answers

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Reaction Rates And Thermochemistry Answers
- 9. Balancing eBooks and Physical Books Reaction Rates And Thermochemistry Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Reaction Rates And Thermochemistry Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Reaction Rates And Thermochemistry Answers
 - Setting Reading Goals Reaction Rates And Thermochemistry Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Reaction Rates And Thermochemistry Answers
 - Fact-Checking eBook Content of Reaction Rates And Thermochemistry Answers
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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