

KEY

Unit 9 Outline- Thermodynamics, Reaction Rates, & Chemical Equilibrium Chapters 17 & 18

Essential Skills/ State Standards:

- 9-1. Describe how temperature and heat flow in terms of the motion of molecules (or atoms). (CSS: 7a)
 9-2. Know that chemical processes & phase changes can either release energy (exothermic) or absorb energy (endothermic). (CSS: 7b, 7c)
 9-3. Solve problems involving heat flow and temperature changes, using known values of specific heat ($q = m \cdot c \cdot \Delta T$) & latent heat of phase change ($q = m \cdot \Delta H$). (CSS: 7d)
 9-4. Know the rate of reaction is the decrease in concentration of reactants or the increase in concentration of products with time & depends on such factors as: concentration, temperature, and pressure. (CSS: 8a, 8b)
 9-5. Know the definition and role of activation energy & catalyst in a chemical reaction. (CSS: 8d*, 8c)
 9-6. Know how to use Le Chatelier's principle to predict the effect of changes in concentration, temperature, and pressure. (CSS: 9a)
 9-7. Know that equilibrium is established when the forward and reverse reaction rates are equal. (CSS: 9b)

Vocabulary

- | | | | |
|----------------|-------------------|--|----------------------------|
| • Condensation | • Deposition | • Calorimeter | • Heat of reaction |
| • Freezing | • Heat | • Joule | • Le Chatelier's principle |
| • Vaporization | • Temperature | • Specific heat | • Equilibrium |
| • Melting | • Thermochemistry | • $q = m \cdot c \cdot \Delta T$ | • Activation Energy |
| • Boiling | • Exothermic | • $q = m \cdot \Delta H_{\text{vaporization}}$ | • Catalyst |
| • Sublimation | • Endothermic | • $q = m \cdot \Delta H_{\text{fusion}}$ | |

Study Guide (Use notes, worksheets, and labs in addition to this study guide)

1. According to the collision theory, what 3 things must happen in order for two molecules to react?
 Must collide with each other, Must collide with enough energy, & in the correct orientation

2. What is the importance of the activation energy in a chemical reaction?
 Amount of energy needed for a reaction to occur.

3. Using the collision theory, explain if the following would increase/ decrease the rate of reaction:
 $2 \text{HCl(aq)} + \text{Mg(s)} \rightarrow \text{MgCl}_2\text{(aq)} + \text{H}_2\text{(g)}$

- a. Increase temp: Increase, molecules collide more
 b. Increase HCl: Increase, more collisions because more reactants to collide
 c. Increase pressure: Hard to say, wouldn't affect the reactants directly because none of them are gases.

One of the products is a gas, so it might cause the gas particles to collide more with the reactants and could slow it down. Tough call.. (this one won't be on the test).

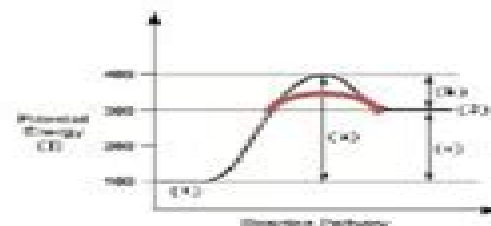
- d. Remove MgCl_2 : Increase, less products means less chance for them to interfere with reactant collisions
 e. Add H_2 : Decrease, more products mean they interfere more with reactants and the reactants would collide less.

4. Identify the following terms with their correct letter from the graphs:

- * Reactants- 1
 * Products- 2
 * Activation energy (number/ letter & amount)-
 A (300 Joules)
 * Is this reaction exothermic or endothermic?

Endothermic

5. Explain how a catalyst affects the activation energy in a chemical reaction. (Also, show on the above graph what a catalyst might do). It lowers the activation energy (see on graph).
 Speeds it up by lowering the activation energy.



Reaction Rates And Equilibrium Test Answers

Clifford Lane



Reaction Rates And Equilibrium Test Answers:

Chemistry Rob Ritchie, 2008-07 Divided into seven manageable day sections this timed revision programme covers essential GCSE topics in double page spreads These spreads indicate how much time should be spent on each section and combine clear and concise explanations flow charts spidergrams and illustrations with progress check questions and answers

Access to Chemistry Avril Higton, Mike Clemmet, Elaine Golding, Alan V Jones, 2007-10-31 This text has been specifically designed to prepare people with previously limited chemical knowledge for entrance into science related courses such as Foundation and Access courses which involve chemistry in higher education Until now there have been no texts available for use on these courses and this book fills that gap Access to Chemistry effectively forms a self study course which is split into separate modules and units covering the full spread of concepts required for those needing a basic knowledge of chemistry The material is presented in a friendly and easy to use manner which allows the student to pace their acquisition of knowledge and gain increasing confidence in order to succeed in understanding essential relevant concepts Other useful features of this book include starter diagnostic tests worked examples and self study tests with answers at the end of each unit In addition to Access or Foundation course students and their tutors to whom this book will prove essential it will have an appeal also as a revision text for those needing a refresher after a break in the subject In addition it will be of interest to members of the general public who wish to better educate themselves on chemical matters as it provides a clear and useful insight into areas such as health home chemicals business market trends and gardening

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explanations It includes 3 full length practice tests with detailed explanations Review of all essential content from chemical equations to kinetics to electron configurations Helpful study lists of key lab equipment and a cheat sheet of important equations Key strategies that will help maximize your score Tons of sample problems and drills with detailed explanations

Chemistry Jerry A. Bell, American Chemical Society, 2005 The American Chemical Society has launched an activities based student centered approach to the general chemistry course a textbook covering all the traditional general chemistry topics but arranged in a molecular context appropriate for biology environmental and engineering students Written by industry chemists and educators Chemistry combines cooperative learning strategies and active learning techniques with a powerful media supplements package to create an effective introductory text Online description Cracking the SAT Chemistry Subject Test Princeton Review, 2014-12-09 EVERYTHING YOU NEED TO HELP SCORE A PERFECT 800 Equip yourself to ace the SAT Chemistry Subject Test with The Princeton Review's comprehensive study guide including 3 full length practice tests thorough reviews of key chemistry topics and targeted strategies for every question type We don't have to tell you how tough SAT Chemistry is or how helpful a stellar exam score can be for your chances of getting into your top choice college Written by the experts at The Princeton Review Cracking the SAT Chemistry Subject Test arms you to take on the test and achieve your highest score Techniques That Actually Work Tried and true strategies to help you avoid traps and beat the test Tips for pacing yourself and guessing logically Essential tactics to help you work smarter not harder Everything You Need to Know for a High Score Expert subject reviews for every test topic Up to date information on the SAT Chemistry Subject Test Score conversion tables for accurate self assessment Practice Your Way to Perfection 3 full length practice tests with detailed answer explanations Hands on experience with all three question types in each content chapter Complete study sheet of core formulas and terms **Chemistry in a Month** John Wilkinson, 2003 *SAT Subject Test Chemistry* Joseph A.

Mascetta, Mark Kernion, 2020-12-01 Note College Board has discontinued the SAT Subject Tests in the US The tests will be available outside the US in June 2021 and then be discontinued Barron's SAT Subject Test Chemistry with 7 Practice Tests features in depth review of all topics on the exam and full length practice tests in the book and online This edition includes One full length diagnostic test to help you assess your strengths and weaknesses Comprehensive review of all topics on the exam including introductory chemistry atomic structure and the periodic table bonding chemical formulas gases and laws stoichiometry liquids solids and phase changes chemical reactions and thermochemistry chemical reactions chemical equilibrium acids bases and salts oxidation reduction carbon and organic chemistry and the laboratory Four full length practice tests that reflect the actual SAT Subject Test Chemistry exam in length question types and degree of difficulty Two full length online practice tests with answer explanations and automated scoring Appendices which include the periodic table important equation constant and data tables and a glossary of chemistry terms **Cracking the AP Chemistry Exam, 2020 Edition** The Princeton Review, 2019-10-22 Make sure you're studying with the most up to date prep materials Look for

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