

Worksheet 1-1

- Measuring Reaction Rates

 Name - _____
 Due Date - _____

Answer the questions in the space provided. You must show all of your work to receive full marks. All answers MUST be rounded to the correct number of Sig. Figures.

1. A chemist wishes to determine the rate of reaction of zinc with hydrochloric acid. The equation for the reaction is:



A piece of zinc is dropped into 1.00 L of 0.100 M HCl and the following data were obtained:

Time	Mass of Zinc
0 s	0.016 g
4 s	0.014 g
8 s	0.012 g
12 s	0.010 g
16 s	0.008 g
20 s	0.006 g

$$\begin{array}{r} 0.016\text{g} \\ - 0.006\text{g} \\ \hline 0.010\text{g} \end{array}$$

- a) Calculate the **Rate of Reaction** in grams of Zn consumed per second.

$$\text{rate} = \frac{\Delta g}{\Delta \text{time}} = \frac{0.010\text{g}}{20\text{sec}}$$

Answer 5.0×10^{-4}

- b) Calculate the **Rate of Reaction** in moles of Zn consumed per second.

$$5.0 \times 10^{-4} \frac{\text{g}}{\text{sec}} \times \frac{1\text{mol}}{65.4\text{g}} =$$

Answer $7.6 \times 10^{-6} \frac{\text{mol}}{\text{sec}}$

- c) Write out the complete ionic equation for the reaction.



- d) What will happen to the $[\text{H}^{+}]$ as the reaction proceeds? decrease

- e) What will happen to the $[\text{Cl}^{-}]$ as the reaction proceeds? nothing, it is a spectator

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Table of Contents Reaction Rate Practice Problems With Answers

1. Understanding the eBook Reaction Rate Practice Problems With Answers
 - The Rise of Digital Reading Reaction Rate Practice Problems With Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Reaction Rate Practice Problems With Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Reaction Rate Practice Problems With Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Reaction Rate Practice Problems With Answers
 - Personalized Recommendations
 - Reaction Rate Practice Problems With Answers User Reviews and Ratings
 - Reaction Rate Practice Problems With Answers and Bestseller Lists
5. Accessing Reaction Rate Practice Problems With Answers Free and Paid eBooks
 - Reaction Rate Practice Problems With Answers Public Domain eBooks
 - Reaction Rate Practice Problems With Answers eBook Subscription Services
 - Reaction Rate Practice Problems With Answers Budget-Friendly Options
6. Navigating Reaction Rate Practice Problems With Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Reaction Rate Practice Problems With Answers Compatibility with Devices
 - Reaction Rate Practice Problems With Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Reaction Rate Practice Problems With Answers
 - Highlighting and Note-Taking Reaction Rate Practice Problems With Answers
 - Interactive Elements Reaction Rate Practice Problems With Answers
8. Staying Engaged with Reaction Rate Practice Problems With Answers

- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Reaction Rate Practice Problems With Answers
9. Balancing eBooks and Physical Books Reaction Rate Practice Problems With Answers
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Reaction Rate Practice Problems With Answers
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Reaction Rate Practice Problems With Answers
- Setting Reading Goals Reaction Rate Practice Problems With Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Reaction Rate Practice Problems With Answers
- Fact-Checking eBook Content of Reaction Rate Practice Problems With Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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