

**CHEMFILE MINI-GUIDE TO PROBLEM SOLVING**

## *The Ideal Gas Law*

In 1811, the Italian chemist Amedeo Avogadro proposed the principle that *equal volumes of gases at the same temperature and pressure contain equal numbers of molecules*. He determined that at standard temperature and pressure, one mole of gas occupies 22.414 10 L (usually rounded to 22.4 L).

At this point, if you know the number of moles of a gas, you can use the molar volume of 22.4 L/mol to calculate the volume that amount of gas would occupy at STP. Then you could use the combined gas law to determine the volume of the gas under any other set of conditions. However, a much simpler way to accomplish the same task is by using the ideal gas law.

The *ideal gas law* is a mathematical relationship that has the conditions of standard temperature (273 K) and pressure (1 atm or 101.3 kPa) plus the molar gas volume (22.4 L/mol) already combined into a single constant. The following equation is the mathematical statement of the ideal gas law.

$$PV = nRT$$

in which

$P$  = the pressure of a sample of gas

$V$  = the volume of a sample of gas

$n$  = the number of moles of gas present

$T$  = the Kelvin temperature of the gas

$R$  = the ideal gas constant, which combines standard conditions and molar volume into a single constant

The value of the ideal gas constant,  $R$ , depends on the units of  $P$  and  $V$  being used in the equation. Temperature is always in kelvins and amount of gas is always in moles. The most common values used for  $R$  are shown below.

Units of $P$ and $V$	Value of $R$
Atmospheres and liters	$0.0821 \frac{\text{L} \cdot \text{atm}}{\text{mol} \cdot \text{K}}$
Kilopascals and liters	$8.314 \frac{\text{L} \cdot \text{kPa}}{\text{mol} \cdot \text{K}}$

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