

## Gas Stoichiometry Worksheet

Name:

Solve all the following gas law problems. Show all work, answers are given at the end of the problem.

### Molar Volume

1. Calculate the number of moles contained in 550.mL of carbon dioxide at STP. (0.0246mol)

2. Calculate the mass of 1.50 L of  $\text{CH}_4$  at STP. (1.07g)

3. Calculate the volume in liters of 50.0 grams of nitrogen dioxide at STP. (24.3L)

### Gas Stoichiometry

4. Given the following reaction:  $\text{NH}_4\text{NO}_2 (\text{s}) \rightarrow \text{N}_2 (\text{g}) + 2\text{H}_2\text{O} (\text{g})$  How many liters of nitrogen gas is produced if 50.0L of water is produced at STP? (25.0L)

5. Given the following reaction:  $\text{CuO} (\text{s}) + \text{H}_2 (\text{g}) \rightarrow \text{Cu} (\text{s}) + \text{H}_2\text{O} (\text{g})$  If 250. L of hydrogen gas are used to reduce copper (II) oxide at STP, what mass of copper is to be expected? (709g)

6. How many grams of solid lithium must be added to liquid water in order to obtain 15.0 L of hydrogen gas at STP? (9.29g)



# Self Check Answers Gas Stoichiometry

**Ralph H. Petrucci**



## Self Check Answers Gas Stoichiometry:

**Basic Chemistry** Steven S Zumdahl,1996      *General Chemistry Workbook* Daniel C. Tofan,2010-07-28 This workbook is a comprehensive collection of solved exercises and problems typical to AP introductory and general chemistry courses as well as blank worksheets containing further practice problems and questions It contains a total of 197 learning objectives grouped in 28 lessons and covering the vast majority of the types of problems that a student will encounter in a typical one year chemistry course It also contains a fully solved 50 question practice test which gives students a good idea of what they might expect on an actual final exam covering the entire material      Introductory Chemistry Steven S. Zumdahl,1993

### **Student's Guide to Introduction to Chemical Principles by Edward I. Peters, 2d Ed** Peter P. Berlow,1978

Basic Concepts of Chemistry, Study Guide Leo J. Malone,2000 The text s three main goals are to introduce chemistry as a living relevant science to encourage learning and critical thinking and to help readers overcome the math difficulties that impede their progress in chemistry Designed to help readers master the principles of general chemistry As a prep book it promotes active involvement with the material There are special features throughout that reinforce concepts and help to develop strong problem solving and study skills Updated to Include an Interactive Learning Ware problems CD containing several of the chapter ending problems from the book in an interactive tutorial with feedback to help readers set up and solve problems      *Study Guide to accompany Basic Concepts of Chemistry, 7th Edition* Leo J. Malone,2003-02-20 Work more effectively and gauge your progress along the way This Study Guide that is designed to accompany Malone s Chemistry 7th Edition includes chapter summaries new terms self tests answers to self tests and solutions to selected problems This easy to read introduction presents chemistry as a living relevant science Chemistry 7th Edition encourages critical thinking and helps readers overcome the math difficulties that often prevent them from developing a full understanding of the subject This new seventh edition builds on its core strengths of pedagogy driving the connections between ideas mathematics in context not just an appendix and an extensive problem solving emphasis with an updated design and more molecular art In addition the seventh edition expands its applications and online options One of the briefest books in the market it still provides sufficient depth for the basic concepts of chemistry      **Chemistry, Student Study Guide** James E. Brady,Fred Senese,2008-01-28 The image on the front cover depicts a carbon nanotube emerging from a glowing plasma of hydrogen and carbon as it forms around particles of a metal catalyst Carbon nanotubes are a recently discovered allotrope of carbon Three other allotropes of carbon buckyballs graphite and diamond are illustrated at the left as is the molecule methane  $\text{CH}_4$  from which nanotubes and buckyballs can be made The element carbon forms an amazing number of compounds with structures that follow from simple methane found in natural gas to the complex macromolecules that serve as the basis of life on our planet The study of chemistry also follows from the simple to the more complex and the strength of this text is that it enables students with varied backgrounds to proceed together to significant levels of achievement      Chemical Principles

Study Guide/Solutions Manual John Krenos, Joseph Potenza, Carl Hoeger, 2007-01-18 Written for general chemistry courses Chemical Principles helps students develop chemical insight by showing the connection between chemical principles and their applications **The Software Encyclopedia 2000** Bowker Editorial Staff, 2000-05 *General Chemistry with*

*Qualitative Analysis* Ralph H. Petrucci, Robert K. Wismer, 1987 **Applications of Spectroscopy, Photochemistry and**

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**Chemical Principles, Fifth Edition, Zumdahl** Paul B. Kelter, Steven S. Zumdahl, 2004-04 NBS Special Publication

, 1968 **Nuclear Science Abstracts** , 1976-02 *Mass Transport in Oxides* J. B. Wachtman, Alan D. Franklin, 1968

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The book delves into Self Check Answers Gas Stoichiometry. Self Check Answers Gas Stoichiometry is a crucial topic that needs to be grasped by everyone, ranging from students and scholars to the general public. The book will furnish comprehensive and in-depth insights into Self Check Answers Gas Stoichiometry, encompassing both the fundamentals and more intricate discussions.

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- Chapter 4: Self Check Answers Gas Stoichiometry in Specific Contexts
- Chapter 5: Conclusion

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This book is crafted in an easy-to-understand language and is complemented by engaging illustrations. It is highly recommended for anyone seeking to gain a comprehensive understanding of Self Check Answers Gas Stoichiometry.

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