

The Periodic Table and Chemical Properties

Textbook pages 52–63

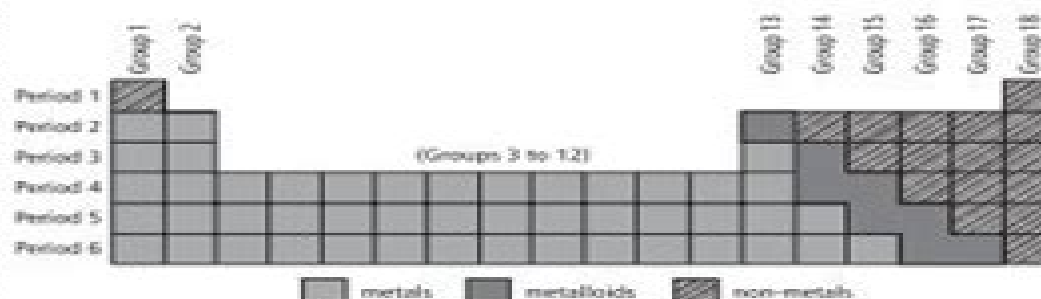
Before You Read

Imagine you had to classify all of the different elements. How would you organize them? How might properties help you decide? Write your thoughts below.

Mark the Text

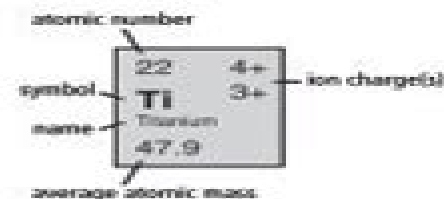
Identify Details

As you skim the section, use one colour to highlight the text or labels that talk about the periodic table. Use another colour to highlight facts about atomic number and other facts about each element.



What is the periodic table?

The **periodic table** places each element on a chart based on its chemical and physical properties. Each box of the periodic table lists information about the atoms that make up an element. You can see:



- ◆ the name of the element
- ◆ the chemical symbol of the element
- ◆ the **atomic number** of the element—the number of protons in the nucleus of each one of its atoms
- ◆ the **average atomic mass** of the element—the weighted average of the masses of the atoms of an element
- ◆ the **ion charge** (or charges) of the element—the electric charge of its atoms when they gain or lose electrons. If the atoms can gain or lose electrons in more than one way, they will have a **multiple ion charge**. Ⓢ

✓ Reading Check

1. What five facts about an element can you find in the periodic table?

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Christian G. Meyer



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Solar Age ,1984 Peter Norton's Computing Fundamentals Peter Norton,1996-11 Presents a fresh approach to computer concepts in a concise 12 chapter text This book is designed for courses that place equal emphasis on computer concepts and hands on learning Its includes an appendix on the ethical considerations of navigating cyberspace It provides an optional CD ROM containing simulations and student activities *Introd to Physical Sci Chapter 3 Atoms Elements and Periodic Table Cr 635ga 02* McGraw-Hill Staff,2001-08-01 **Integrated Physics and Chemistry, Chapter 3 Activities**
Paradigm Accelerated Curriculum,2005-01-01 Key topics Chemical nomenclature Lavoisiers list of elements sulfur diamonds graphite coal medieval metals platinum zinc cobalt nickel manganese molybdenum tungsten gases in the atmosphere air pressure and humidity Henry Cavendish hydrogen nitrogen fertilizers and explosives dynamite laughing gas IPC consists of twelve chapters of text and twelve companion student activity books This course introduces students to the people places and principles of physics and chemistry It is written by internationally respected scientist author John Hudson Tiner who applies the vignette approach which effectively draws readers into the text and holds attention The author and editors have deliberately avoided complex mathematical equations in order to entice students into high school level science Focus is on the people who contributed to development of the Periodic Table of the Elements Students learn to read and apply the Table while gaining insight into basic chemistry and physics This is one of our most popular courses among high school students especially those who have a history of under performance in science courses due to poor mathematical and reading comprehension skills The course is designed for two high school transcript credits Teachers may require students to complete all twelve chapters for two transcript credits or may select only six chapters to be completed for one transcript credit for Physical Science Physics or Chemistry Compliance with state and local academic essential elements should be considered when specific chapters are selected by teachers As applicable to local policies transcript credit may be assigned as follows when students complete all 12 chapters Physical Science for one credit and Chemistry for one credit or Integrated Physics and Chemistry for two credits May require supplemental local classes labs

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In an electronic digital era where connections and knowledge reign supreme, the enchanting power of language has become more apparent than ever. Its power to stir emotions, provoke thought, and instigate transformation is really remarkable. This extraordinary book, aptly titled "**Section 3 The Periodic Table Mcgraw Hill Worksheet Answers**," published by a very acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound effect on our existence. Throughout this critique, we will delve into the book's central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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