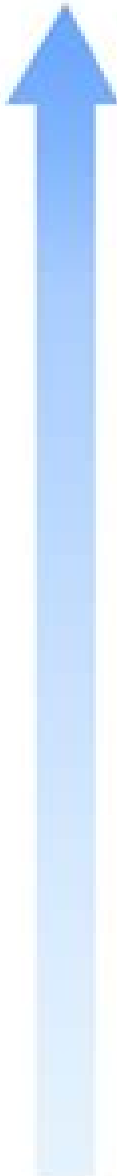


Reactivity Series

Metal (Ion)	Reactivity	Most reactive
Potassium (K ⁺)	Reacts with water	
Sodium (Na ⁺)		
Lithium (Li ⁺)		
Barium (Ba ²⁺)		
Strontium (Sr ²⁺)		
Calcium (Ca ²⁺)		
Magnesium (Mg ²⁺)	Reacts with acid	
Aluminum (Al ³⁺)		
Manganese (Mn ⁺)		
Zinc (Zn ²⁺)		
Chromium (Cr ²⁺)		
Iron (Fe ²⁺)		
Cadmium (Cd ²⁺)		
Cobalt (Co ²⁺)		
Nickle (Ni ²⁺)		
Tin (Sn ²⁺)		
Lead (Pb ²⁺)		
Hydrogen (H ⁺)	Included for comparison	
Antimony (Sb ²⁺)	Highly unreactive	
Bismuth (Bi ²⁺)		
Copper (Cu ²⁺)		
Mercury (Hg ²⁺)		
Silver (Ag ⁺)		
Gold (Au ²⁺)		
Platinum (Pt ⁺)	Least reactive	

Reactivity Series Pogil

John C. Kotz, Keith F. Purcell



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Reactivity Series & Displacement Reactions, 2006 Reactivity Series, 2006 Chemical Reactions, Materials and Particles Gill Murphy, 2006-03 Specials are a teaching resource designed for KS3 students whose literacy skills are considerably lower than their age. These books have an older format to counteract the simple text and cover various topics. They include activities, visuals, and assessment sheets as well as teacher pages. **Mechanism and Reactivity for a Series**

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Approaches and Perspectives D. Heidrich, 2013-01-07. The so-called reaction path (RP) with respect to the potential energy or the Gibbs energy/free enthalpy is one of the most fundamental concepts in chemistry. It significantly helps to display and visualize the results of the complex microscopic processes forming a chemical reaction. This concept is an implicit component of conventional transition state theory (TST). The model of the reaction path and the TST form a qualitative framework which provides chemists with a better understanding of chemical reactions and stirs their imagination. However, an exact calculation of the RP and its neighbourhood becomes important when the RP is used as a tool for a detailed exploring of reaction mechanisms and particularly when it is used as a basis for reaction rate theories above and beyond TST. The RP is a theoretical instrument that now forms the theoretical heart of direct dynamics. It is particularly useful for the interpretation

of reactions in common chemical systems A suitable definition of the RP of potential energy surfaces is necessary to ensure that the reaction theories based on it will possess sufficiently high quality Thus we have to consider three important fields of research Analysis of potential energy surfaces and the definition and best calculation of the RPs or at least of a number of selected and chemically interesting points on it The further development of concrete versions of reaction theory beyond TST which are applicable for common chemical systems using the RP concept Chemistry & Chemical Reactivity John C.

Kotz, Keith F. Purcell, 1987 **Reactivity Patterns in a Series of Novel Monothiocarbamates, Dithiocarbamates, and Sulfur-rich Dithiolates** Donald M. Baird, 1982 **Chemistry and Chemical Reactivity** John C. Kotz, Keith F. Purcell, 1987

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- Chapter 5: Conclusion

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