

Worksheet - Reaction RatesUse this reaction for the questions below: $\text{C}_6\text{H}_{12}\text{O}_6 (\text{s}) + 6 \text{O}_2 (\text{g}) \rightarrow 6 \text{H}_2\text{O} (\text{g}) + 6 \text{CO}_2 (\text{g})$

- What happens to the concentrations of:
 - $\text{C}_6\text{H}_{12}\text{O}_6$ & O_2 as the reaction proceeds $\rightarrow$?
 - H_2O + CO_2 as the reaction proceeds $\rightarrow$?
- According to the collision theory, what 3 circumstances are needed for $\text{C}_6\text{H}_{12}\text{O}_6$ & O_2 to react?
- What is the activation energy for a chemical reaction?
- Use the equation & the collision theory to explain:



Change in condition:	Does this <u>increase</u> or <u>decrease</u> the rate of reaction?	Explain why
a. <u>Increasing the temperature</u>	Ex: Increase (spreads up)	Ex: Molecules move faster & collide more = increased rxn rate
b. <u>Increasing the concentration</u> of $\text{C}_6\text{H}_{12}\text{O}_6$		
c. <u>Decreasing the concentration</u> of O_2		
d. <u>Increase the surface area</u> by chewing up food in your mouth.		
e. <u>Decreasing the temperature</u>		
f. <u>Increasing the pressure</u> in the container		
g. <u>Decreasing the concentration</u> of H_2O		
h. <u>Increasing the volume</u> of the container the reaction occurs in		
i. <u>Increasing the concentration</u> of CO_2		
j. <u>Using a catalyst</u> (like salivary amylase)		

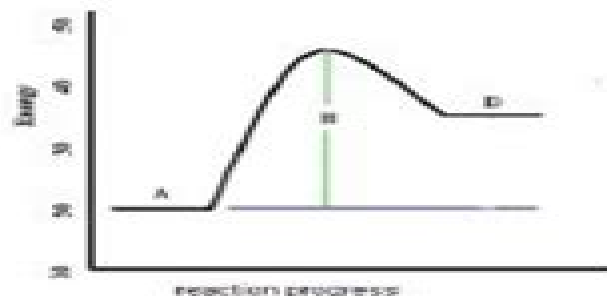
- On the accompanying energy diagram, label the following terms:

a. reactants b. products c. activation energy

- On the graph to the right, draw and label what this diagram would look like if a catalyst was added to the reaction.

- Graph reading

- How much energy (kJ) do the reactants have?
- How much energy (kJ) do the products have?
- How much energy (kJ) is required to activate this un-catalyzed reaction?
- Is this reaction endothermic or exothermic? How do you know?



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Peipei Pang



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