

CHAPTER 16 REVIEW***Reaction Energy*****SECTION 2****SHORT ANSWER** Answer the following questions in the space provided.

1. For the following examples, state whether the change in entropy favors the forward or reverse reaction:

forward reactionreverse reactionforward reactionreverse reaction

2. (change in free energy) $\Delta G =$ (change in enthalpy) $\Delta H - T$ (change in entropy) ΔS

negative

a. Write an equation that shows the relationship between enthalpy, ΔH , entropy, ΔS , and free energy, ΔG .

b. For a reaction to occur spontaneously, the sign of ΔG should be _____

3. Consider the following equation: $\text{NH}_3(g) + \text{H}_2\text{O}(l) \rightleftharpoons \text{NH}_4^+(aq) + \text{OH}^-(aq) + \text{energy}$

true

a. The enthalpy factor favors the forward reaction. True or False?

reverse reaction

b. The sign of $T\Delta S^\circ$ is negative. This means the entropy factor favors the _____

- c. Given that ΔG° for the above reaction in the forward direction is positive, which term is greater in magnitude and therefore predominates, $T\Delta S$ or ΔH ?

$T(\text{change in } S)$ is greater than $(\text{change in } H)$; $(\text{change in } H)$ is negative, but $[\Delta G]$ is positive, indicating the reaction is not spontaneous. The randomness of the reactants and the temperature of the reaction overshadow the exothermic factor.

4. Consider the following equation for the vaporization of water:

endothermic

a. Is the forward reaction exothermic or endothermic?

reverse reaction

b. Does the enthalpy factor favor the forward or reverse reaction?

forward reaction

c. Does the entropy factor favor the forward or reverse reaction?

Reaction Energy Section Answer Key

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