

THERMOCHEMISTRY ASSIGNMENT

1. Thermodynamic stability of graphite is defined as the heat required for conversion of graphite into diamond. The enthalpy change for this reaction is, $\Delta H = 1.9$ kJ/mole. Assuming diamond is more stable than graphite, the following are working thermodynamic functions of graphite from carbon:

The above reaction is exothermic and as such will release energy and increase the temperature of the hot water the graphite will be being produced. As a result the hot water can be used to generate electricity with various types of turbines.

2. Thermodynamic heat capacity of diamond is 0.005 J/g°C. Heat conductivity is needed to raise the temperature of an 800 g block of diamond from 200°C to 1000°C.

(a) Energy

$$Q = mc\Delta T = 800 \times 0.005 \times (1000 - 200) = 2000 \text{ J}$$

(b) Temperature

$$T = \frac{Q}{mc} = \frac{2000}{800 \times 0.005} = 500^\circ\text{C}$$

3. Thermodynamic heat capacity of graphite is 0.075 J/g°C. Maximum thermodynamic temperature for water the temperature of 100 g of graphite is 100°C.

(a) Energy

$$Q = mc\Delta T = 100 \times 0.075 \times (100 - 0) = 7.5 \text{ J}$$

(b) Temperature

$$T = \frac{Q}{mc} = \frac{7.5}{100 \times 0.075} = 100^\circ\text{C}$$

- (c) Water the temperature of 100 g of graphite from 200 to 1000 K.

$$Q = mc\Delta T = 100 \times 0.075 \times (1000 - 200) = 600 \text{ J}$$

$$T = \frac{Q}{mc} = \frac{600}{100 \times 0.075} = 800^\circ\text{C}$$

(d) Energy

$$Q = mc\Delta T = 100 \times 0.075 \times (1000 - 0) = 7500 \text{ J}$$

$$T = \frac{Q}{mc} = \frac{7500}{100 \times 0.075} = 1000^\circ\text{C}$$

(e) The water

4. To raise 750 g of water the temperature of 100 g of lead by 100 K. What is thermodynamic heat capacity of lead?

$$Q = mc\Delta T$$

$$7500 = 750 \times c \times 100$$

$$c = \frac{7500}{750 \times 100}$$

$$c = 0.01 \text{ J/g°C}$$

$$T = \frac{Q}{mc} = \frac{7500}{750 \times 0.01} = 1000^\circ\text{C}$$

5. A substance is initially contained 100 g of water at a temperature of 100°C. 100 g of graphite is added. After 100 g of 100°C is added to the water, the final temperature is 100°C. 100 g of graphite is added to the water. The final temperature is 100°C. The final temperature of the water is 100°C.

Reaction Rates And Thermochemistry Answers

Arshad Iqbal



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substances for lives However we didn't have the clear answer to the following questions until the 1920s why molecules can exist in stable as rigid networks between atoms and why molecules can change into different types of molecules The most important event for solving the puzzles is the discovery of the quantum mechanics Quantum mechanics is the theory for small particles such as electrons and nuclei and was applied to hydrogen molecule by Heitler and London at 1927 The pioneering work led to the clear explanation of the chemical bonding between the hydrogen atoms This is the beginning of the quantum chemistry Since then quantum chemistry has been an important theory for the understanding of molecular properties such as stability reactivity and applicability for devices This book is devoted for the theoretical foundations and innovative applications in quantum chemistry

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