

CHEMISTRY 1211 TEST #3

Use a Scantron sheet for Multiple Choices (3pts each). Do not write on the Periodic Table.

Use back of this page to show your work/answer for the Big Question.

1mi=1.609km; 1lbs=453.6g; 1gal=3.785L; 1mL=1cm³; 1=10⁻¹⁰⁰ m;
1cal=4.184J; g=9.807m/s
1atm=760mmHg; Density(Hg)=13.60g/mL; c=2.9979x10⁸m/s; h=6.626x10⁻³⁴J R_∞=2.180x10⁻¹⁸J;
m_e=9.109x10⁻³¹ kg; R=8.315J/k mol.

1. Which photon has the lowest energy ?

A(X-ray), B(Ultraviolet radiation, UV), C(Visible light), D(Infrared radiation, IR), **E(Microwave)**.

2. Energy of an electromagnetic wave increases with (a) increasing, (b) decreasing, frequency; therefore a blue light has a higher energy than a red light because of its (c) higher, (d) lower frequency than that of a red light. **A(a,c)**, B(a,d), C(b,c), D(b,d).

3. Frequency of a radio wave is 104.7 MHz. It's wavelength must be about A(3 fm), **B(3 cm)**, C(3 m), D(3 km), E(3 Mm).

4. Find energy of an X-ray with a wavelength of 1.0x10⁻¹⁰ m (= 0.10 nm). Report it (a) per photon and (b) per mole.

A(1.986x10⁻¹⁷/photon; 1.196x10⁶kJ/mol), B(1.986x10⁻¹²J/photon; 1.196x10⁶kJ/mol), C(1.986x10⁻¹³J/photon; 1.196x10⁶kJ/mol), D(1.986x10⁻¹²J/photon; 1.196x10⁶kJ/mol), **E(1.986x10⁻¹⁷J/photon; 1.196x10⁶kJ/mol)**.

5. Who got the Nobel Prize in 1921 with an idea that light behave as if it is a particle whose energy is directly proportional to its frequency ?

A(Balmer), B(Planck), **C(Einstein)**, D(Bohr), E(Heisenberg).

6. What is the energy (J) of an electron at the third energy level (i.e., in the L shell with the principle quantum number n=3) in a H atom?

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